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## PACKET #1

## Unit 1: Chemistry of Life, Part I

*Reading: BSCS Text Chapter 1, 2.9, 2.10**Learning Objectives:**Topic 1: Basic Chemistry and Bonding*

1. Explain the relationships among atoms and molecules (1.1).
2. Describe and diagram the structure of an atom, including protons, neutrons, electrons (1.2).
3. Use atomic number on the periodic table to determine number of protons and electrons in an atom (class).
4. Explain why atoms form covalent bonds and use a Bohr model/HONC rule to predict bonding behavior(1.4).
5. Define and draw an ion, and provide examples (1.4).
6. Explain why atoms form ionic bonds. Given the number of valence electrons in an atom, predict whether the atom is likely to participate in ionic bonding.

*Topic 2: Properties of water*

7. Explain why water molecules are considered polar and how this helps hydrogen "bonds" form (1.4)
8. How is an **intermolecular** force (like a hydrogen "bond") different from a **intramolecular** force (like a covalent bond)?
9. *Explain* the following properties of water (including how these are biologically relevant): high surface tension, high specific heat, solid water is less dense than liquid water, water as the "universal solvent," and capillary action. Along the way, be sure to address cohesion and adhesion (p. 189, teacher notes).
10. Differentiate between acids and bases on a pH scale.

**Key Terms/Concepts**

Molecules	Covalent Bond	Intermolecular force
Atoms	Ionic Bond	Intramolecular force
Element	Ion	Cohesion
Subatomic particles	Ionization	Adhesion
Proton	Hydrogen bond	Capillary action
Neutron	Polar molecule	Surface tension
Electron	Atomic number	Specific heat
Solution	Solute	Salt
Acid	Base	Solvent

# Grade 1chemistry Unit Practice Test

**Nebraska. University,University of  
Nebraska (Lincoln campus)**

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### **Table of Contents Grade 1chemistry Unit Practice Test**

1. Understanding the eBook Grade 1chemistry Unit Practice Test
  - The Rise of Digital Reading Grade 1chemistry Unit Practice Test
  - Advantages of eBooks Over Traditional Books
2. Identifying Grade 1chemistry Unit Practice Test
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Grade 1chemistry Unit Practice Test
  - User-Friendly Interface
4. Exploring eBook Recommendations from Grade 1chemistry Unit Practice Test
  - Personalized Recommendations
  - Grade 1chemistry Unit Practice Test User Reviews and Ratings

- Grade 1chemistry Unit Practice Test and Bestseller Lists
- 5. Accessing Grade 1chemistry Unit Practice Test Free and Paid eBooks
  - Grade 1chemistry Unit Practice Test Public Domain eBooks
  - Grade 1chemistry Unit Practice Test eBook Subscription Services
  - Grade 1chemistry Unit Practice Test Budget-Friendly Options
- 6. Navigating Grade 1chemistry Unit Practice Test eBook Formats
  - ePub, PDF, MOBI, and More
  - Grade 1chemistry Unit Practice Test Compatibility with Devices
  - Grade 1chemistry Unit Practice Test Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Grade 1chemistry Unit Practice Test
  - Highlighting and Note-Taking Grade 1chemistry Unit Practice Test
  - Interactive Elements Grade 1chemistry Unit Practice Test
- 8. Staying Engaged with Grade 1chemistry Unit Practice Test
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Grade 1chemistry Unit Practice Test
- 9. Balancing eBooks and Physical Books Grade 1chemistry Unit Practice Test
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Grade 1chemistry Unit Practice Test
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Grade 1chemistry Unit Practice Test
  - Setting Reading Goals Grade 1chemistry Unit Practice Test
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Grade 1chemistry Unit Practice Test
  - Fact-Checking eBook Content of Grade 1chemistry Unit Practice Test
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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