

Half Modern Chemistry Review

CHAPTER 5: THE PERIODIC TABLE

The following pages contain the best (but not all) of the information for this chapter in class.

Focus on this content, but make sure to review other topics, activities, handouts, problems, etc.

If you already take advanced-level SCIENCE class, you should at least be able to PASS this test.

**** Test items will be recall, examples, and/or application of this content. ****

COURSES:

- College-level (with passing or substandard/intermediate success) SC 121
- Community-college-level (with or without actual passing) SC 121
- 5.1: Describe the organization of the periodic table; distinguish between periods and groups; IT & AP
- 5.2: Use the Periodic Table to identify elements about their properties (chemical and physical), classification, energy, and atomic size; of that elements; IT & APC
- 5.3: Identify common nonmetals, metalloids, and metal groups; identify the representative groups and/or main groups (groups 1, 2, 13, 14, 15, 16, 17, 18) IT & AP
- 5.4: Compare periodic trends of atomic size, ionization energy, and electronegativity; IT & AP

5.1: PERIODS OF THE PERIODIC TABLE

• Introductory

- periodic law - the law that states that the repeating chemical and physical properties of elements change periodically with the atomic numbers of the elements
- periodic table - an arrangement of the elements in order of their atomic numbers such that elements with similar properties fall in the same column.
- horizontalities - a number of the same number of elements, known as the number, range from 2 (hydrogen) to 18 (barium)
- verticality - set of the same of heavy metals that elements that decrease from hydrogen (atomic number 1) through barium (atomic number 138) on the periodic table

• Chapter Highlights

- This periodic law states that the physical and chemical properties of the elements and periodic functions of these elements, vary.
- This periodic table is an arrangement of the elements in order of their atomic numbers so that elements with similar properties fall in the same column/group/family.
- The columns in the periodic table are referred to as groups.

5.2: ELECTRON CONFIGURATIONS AND THE PERIODIC TABLE

• Introductory

- shell model - one of the elements of Group 1 of the periodic table (lithium, sodium, potassium, rubidium, cesium, and francium)
- shellness model - one of the elements of Group 2 of the periodic table (beryllium, magnesium, calcium, strontium, barium, and radium)
- transition elements - one of the metals that can use the inner shell electrons along the outer shell to bond.
- metalloid elements - an element in the d-block or p-block of the periodic table
- halogens - one of the elements of Group 17 (fluorine, chlorine, bromine, iodine, and astatine); halogens combine with most metals to form salts.

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Table of Contents Holt Chemistry Chapter 5 Review Answers

1. Understanding the eBook Holt Chemistry Chapter 5 Review Answers
 - The Rise of Digital Reading Holt Chemistry Chapter 5 Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Holt Chemistry Chapter 5 Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Holt Chemistry Chapter 5 Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Holt Chemistry Chapter 5 Review Answers

- Personalized Recommendations
 - Holt Chemistry Chapter 5 Review Answers User Reviews and Ratings
 - Holt Chemistry Chapter 5 Review Answers and Bestseller Lists
5. Accessing Holt Chemistry Chapter 5 Review Answers Free and Paid eBooks
 - Holt Chemistry Chapter 5 Review Answers Public Domain eBooks
 - Holt Chemistry Chapter 5 Review Answers eBook Subscription Services
 - Holt Chemistry Chapter 5 Review Answers Budget-Friendly Options
 6. Navigating Holt Chemistry Chapter 5 Review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Holt Chemistry Chapter 5 Review Answers Compatibility with Devices
 - Holt Chemistry Chapter 5 Review Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Holt Chemistry Chapter 5 Review Answers
 - Highlighting and Note-Taking Holt Chemistry Chapter 5 Review Answers
 - Interactive Elements Holt Chemistry Chapter 5 Review Answers
 8. Staying Engaged with Holt Chemistry Chapter 5 Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Holt Chemistry Chapter 5 Review Answers
 9. Balancing eBooks and Physical Books Holt Chemistry Chapter 5 Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Holt Chemistry Chapter 5 Review Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Holt Chemistry Chapter 5 Review Answers
 - Setting Reading Goals Holt Chemistry Chapter 5 Review Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Holt Chemistry Chapter 5 Review Answers

- Fact-Checking eBook Content of Holt Chemistry Chapter 5 Review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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