

## Assessment

**Chapter Test B****Teacher Notes and Answers****8** *Chemical Equations and Reactions***TEST B**

- b
- a
- b
- d
- a
- b
- d
- a
- coefficient
- liquid
- reversible reaction
- $A + X \rightarrow AX$
- $AX \rightarrow A + X$
- $A + BX \rightarrow AX + B$
- $AX + BY \rightarrow AY + BX$
- metal hydroxides
- combustion reaction
- electrolysis
- activity series
- chemical equation
- conservation of mass
- gaseous
- heated
- A word equation and a formula equation are both qualitative. Although they both show the reactants and products of a chemical reaction, they give no information about amounts of reactants or products. A chemical equation, however, does show the relative amounts of the reactants and products in a chemical reaction.
- Three indications of a chemical reaction are the radiation of heat and light, the production of a gas, and the formation of a precipitate.
- zinc sulfide + oxygen  $\rightarrow$   
zinc oxide + sulfur dioxide
- sodium iodide + chlorine  $\rightarrow$   
sodium chloride + iodine
- $Mg + O_2 \rightarrow MgO$
- $CaCO_3 \rightarrow CaO + CO_2$
- decomposition reaction
- single-displacement reaction
- synthesis reaction
- double-displacement reaction
- $2AgNO_3 + CuCl_2 \rightarrow 2AgCl + Cu(NO_3)_2$
- $2PbO_2 \rightarrow 2PbO + O_2$
- $Zn(OH)_2 + 2CH_3COOH \rightarrow$   
 $Zn(CH_3COO)_2 + 2H_2O$
- $Mg(s) + 2H_2O(g) \rightarrow Mg(OH)_2(aq) + H_2(g)$
- no reaction
- $Cl_2(g) + MgBr_2(aq) \rightarrow MgCl_2(aq) + Br_2(l)$
- $Zn(s) + 2HCl(aq) \rightarrow ZnCl_2(aq) + H_2(g)$
- no reaction
- $H_2(s) + CuCl_2(aq) \rightarrow H_2Cl_2(aq) + Cu(s)$
- no reaction
- $Mg(s) + Co(NO_3)_2(aq) \rightarrow$   
 $Mg(NO_3)_2(aq) + Co(s)$

# Holt Chemistry Chapter 8 Test

**Y Pai**



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