

# ICE Practice Problems

#1 – Relatively easy, no ICE table required because eq'm concentrations are given

For the reaction  $\text{CH}_4(\text{g}) + \text{H}_2\text{O}(\text{g}) \leftrightarrow \text{CO}(\text{g}) + 3\text{H}_2(\text{g})$  @  $1500^\circ\text{C}$  an equilibrium mixture of these gases was found to have the following concentrations  $[\text{CO}] = 0.300\text{M}$ ,  $[\text{H}_2] = 0.800\text{M}$  and  $[\text{CH}_4] = 0.400\text{M}$ .  $K_c$  @  $1500^\circ\text{C} = 5.67$ . Determine the equilibrium concentration of  $\text{H}_2\text{O}$  in this mixture.

#2 – Requires an ICE table because you do not know the equilibrium concentrations- no product is yet formed

For the reaction  $\text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \leftrightarrow \text{CO}(\text{g}) + \text{H}_2(\text{g})$  calculate the equilibrium concentrations of all species if 1.000 mol of each reactant is mixed in a 1.000L flask.  $K_c = 5.10$  at the temperature of this reaction.

		$\text{CO}_2(\text{g})$	+	$\text{H}_2\text{O}(\text{g})$	$\leftrightarrow$	$\text{CO}(\text{g})$	+	$\text{H}_2(\text{g})$
[Initial]	I	1.000		1.000		0		0
[Change in]	C							
[Equilibrium]	E							

#3 – Requires an ICE table because you do not know the equilibrium concentrations. The initial concentrations must be calculated- no product is yet formed

For the reaction  $\text{H}_2(\text{g}) + \text{F}_2(\text{g}) \leftrightarrow 2\text{HF}(\text{g})$  calculate the equilibrium concentrations of all species if 3.000 mol of each reactant was added 1.500L flask.  $K_c$  at the temperature of the reaction is  $1.15 \times 10^2$ .

		$\text{H}_2(\text{g})$	+	$\text{F}_2(\text{g})$	$\leftrightarrow$	$2\text{HF}(\text{g})$
[Initial]	I	2.00		2.00		0
[Change in]	C					
[Equilibrium]	E					

#4 – Requires an ICE table because you do not know the equilibrium concentrations. Initial concentrations of reactants are given.

0.200mol of  $\text{H}_2$ , 0.200mol of  $\text{I}_2$ , and 0.200mol of  $\text{HI}$  were placed in a 1.00 L flask and allowed to come to equilibrium. The  $K_c$  value of the reaction at this temperature is 49.5. Determine the equilibrium concentrations of all species.

		$\text{H}_2(\text{g})$	+	$\text{I}_2(\text{g})$	$\leftrightarrow$	$2\text{HI}(\text{g})$
[Initial]	I					
[Change in]	C					
[Equilibrium]	E					

#5 – Requires an ICE table because you do not know the equilibrium concentrations. Initial concentrations of reactants must be calculated and no product is yet formed

For the reaction  $\text{H}_2(\text{g}) + \text{F}_2(\text{g}) \leftrightarrow 2\text{HF}(\text{g})$  calculate the equilibrium concentrations of each species if 3.000 mol of  $\text{H}_2$  and 6.000mol of  $\text{F}_2$  are mixed in a 3.000L flask.  $K_c$  at this temperature is  $1.15 \times 10^2$ .

		$\text{H}_2(\text{g})$	+	$\text{F}_2(\text{g})$	$\leftrightarrow$	$2\text{HF}(\text{g})$
[Initial]	I					
[Change in]	C					
[Equilibrium]	E					

# Ice Practice Problems Answer Key

**E Durkheim**



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## **Table of Contents Ice Practice Problems Answer Key**

1. Understanding the eBook Ice Practice Problems Answer Key
  - The Rise of Digital Reading Ice Practice Problems Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Ice Practice Problems Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ice Practice Problems Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ice Practice Problems Answer Key
  - Personalized Recommendations
  - Ice Practice Problems Answer Key User Reviews and Ratings
  - Ice Practice Problems Answer Key and Bestseller Lists
5. Accessing Ice Practice Problems Answer Key Free and Paid eBooks
  - Ice Practice Problems Answer Key Public Domain eBooks
  - Ice Practice Problems Answer Key eBook Subscription Services
  - Ice Practice Problems Answer Key Budget-Friendly Options
6. Navigating Ice Practice Problems Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Ice Practice Problems Answer Key Compatibility with Devices
  - Ice Practice Problems Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ice Practice Problems Answer Key
  - Highlighting and Note-Taking Ice Practice Problems Answer Key
  - Interactive Elements Ice Practice Problems Answer Key
8. Staying Engaged with Ice Practice Problems Answer Key

- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ice Practice Problems Answer Key
9. Balancing eBooks and Physical Books Ice Practice Problems Answer Key
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ice Practice Problems Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Ice Practice Problems Answer Key
- Setting Reading Goals Ice Practice Problems Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ice Practice Problems Answer Key
- Fact-Checking eBook Content of Ice Practice Problems Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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