

## Released AP Multiple Choice Questions – Acids &amp; Bases / Aqueous Equilibrium

D	33. The pH of 0.1-molar ammonia is approximately (A) 1 (B) 4 (C) 7 (D) 11 (E) 14
E	48. Which of the following ions is the strongest Lewis acid? (A) $\text{Na}^+$ (B) $\text{Cl}^-$ (C) $\text{CH}_3\text{COO}^-$ (D) $\text{Mg}^{2+}$ (E) $\text{Al}^{3+}$
C	49. Each of the following can act as both a Brønsted acid and a Brønsted base EXCEPT (A) $\text{HCO}_3^-$ (B) $\text{H}_2\text{PO}_4^-$ (C) $\text{NH}_4^+$ (D) $\text{H}_2\text{O}$ (E) $\text{HS}^-$
B	53. Which, if any, of the following species is in the greatest concentration in a 0.100-molar solution of $\text{H}_2\text{SO}_4$ in water? (A) $\text{H}_2\text{SO}_4$ molecules (B) $\text{H}_3\text{O}^+$ ions (C) $\text{HSO}_4^-$ ions (D) $\text{SO}_4^{2-}$ ions (E) All species are in equilibrium and therefore have the same concentrations.
E	71. Which of the following reactions does NOT proceed significantly to the right in aqueous solutions? (A) $\text{H}_3\text{O}^+ + \text{OH}^- \rightarrow 2 \text{H}_2\text{O}$ (B) $\text{HCN} + \text{OH}^- \rightarrow \text{H}_2\text{O} + \text{CN}^-$ (C) $\text{Cu}(\text{H}_2\text{O})_6^{2+} + 4 \text{NH}_3 \rightarrow \text{Cu}(\text{NH}_3)_4^{2+} + 4\text{H}_2\text{O}$ (D) $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{HSO}_4^-$ (E) $\text{H}_2\text{O} + \text{HSO}_4^- \rightarrow \text{H}_2\text{SO}_4 + \text{OH}^-$
E	75. If the acid dissociation constant, $K_a$ , for an acid HA is $8 \times 10^{-4}$ at $25^\circ\text{C}$ , what percent of the acid is dissociated in a 0.50-molar solution of HA at $25^\circ\text{C}$ ? (A) 0.08% (B) 0.2% (C) 1% (D) 2% (E) 4%
A	34. All of the following species can function as Brønsted-Lowry bases in solution EXCEPT (A) $\text{H}_2\text{O}$ (B) $\text{NH}_3$ (C) $\text{S}^{2-}$ (D) $\text{NH}_4^+$ (E) $\text{HCO}_3^-$

# General Equilibrium Ap Chemistry Multiple Choice

**N Colangelo**



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In some sort of eaten by displays and the ceaseless chatter of instantaneous conversation, the melodic splendor and mental symphony produced by the published term often fade in to the back ground, eclipsed by the constant noise and distractions that permeate our lives. Nevertheless, located within the pages of **General Equilibrium Ap Chemistry Multiple Choice** a stunning fictional value brimming with organic emotions, lies an immersive symphony waiting to be embraced. Crafted by a masterful musician of language, this fascinating masterpiece conducts readers on a mental journey, well unraveling the concealed melodies and profound affect resonating within each carefully crafted phrase. Within the depths of the moving examination, we can investigate the book is key harmonies, analyze its enthralling writing fashion, and submit ourselves to the profound resonance that echoes in the depths of readers souls.

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