

NOTE:

- If a candidate answers a question/vraag TWICE, only mark the FIRST attempt.
- Consistent accuracy applies in all aspects of the marking memorandum.

LET WEL:

- Indien 'n kandidaat 'n vraag TWEE keer beantwoord, merk slegs die EERSTE poging.
- Volgehoue akkuraatheid is DEURGAANS in ALLE aspekte van die memorandum van toepassing.

QUESTION/VRAAG 1

1.1.1	$3x^2 - 4x = 0$ $x(3x - 4) = 0$ $x = \frac{4}{3} \text{ or } x = 0$	✓ factors ✓ both answers (2)
1.1.2	$x - 6 + \frac{2}{x} = 0$ $x^2 - 6x + 2 = 0$ $x = \frac{-(-6) \pm \sqrt{(-6)^2 - 4(1)(2)}}{2(1)}$ $= \frac{6 \pm \sqrt{28}}{2}$ $x = 0,35 \text{ or } x = 5,65$ <p>OR</p> $x - 6 + \frac{2}{x} = 0$ $x^2 - 6x + 2 = 0$ $(x - 3)^2 = -2 + 9$ $(x - 3) = \pm \sqrt{7}$ $x = 3 \pm \sqrt{7}$ $x = 0,35 \text{ or } x = 5,65$	$\checkmark x^2 - 6x + 2 = 0$ $\checkmark \text{ subs into}$ correct formula $\checkmark x = 0,35$ $\checkmark x = 5,65$ (4)
1.1.3	$x^{\frac{2}{3}} = 4; x > 0$ $x = (2^2)^{\frac{3}{2}}$ $x = 8$ <p>OR</p>	$\checkmark x = (2^2)^{\frac{3}{2}}$ $\checkmark x = 8$ (2)

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web indicate each person accurately note unaffected means normal skin color while affected means the person was blue the

carrier symbols mean the person carried a blue allele but didn't show signs of it on your pedigree chart you must have the following clearly labeled names when provided symbols associated with male or female

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web a 5 5 page activity titled the blue people of troublesome creek focuses on the blue fugates of kentucky this family

suffered from a rare genetic disorder known as methemoglobinemia this is really fascinating and my students love hearing about this

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web deficiency of nadh diaphorase he went out into the hills and found other blue people including zach fugate and his aunt bessie fugate all living near a nearly dead mining town known as hardburly the doctor had noticed that the blue people were not particularly happy about being blue

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web the blue people of troublesome creek a classroom favorite 5 pages of pedigree fun plus a 12 slide powerpoint all about the methemoglobinemia pedigree quiz 10 questions that include mostly autosomal pedigrees

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web feb 21 2012 the most detailed account blue people of troublesome creek was published in 1982 by the university of indiana s cathy trost who described benjy s skin as almost purple the fugate progeny had a genetic condition called methemoglobinemia which was passed down through a recessive gene and blossomed through intermarriage

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web in a nutshell the blue is caused by a hereditary condition called methemoglobinemia it is a rare and harmless genetic disorder in which the body does not make enough of the enzyme diaphorase or any those with methemoglobinemia still have plenty of hemoglobin so they function normally

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web nov 25 2012 martin fugate was an orphan who moved to kentucky in 1820 to find a place to live near troublesome creek family stories say that martin himself was blue the odds were very much against it but martin fugate managed to find and marry a woman who carried the same recessive gene

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web jan 26 2016 the solubility product or ksp is a way of showing how effective a chemical is at dissolving into an aqueous solution the higher the ksp the more the chemical will dissolve into that solvent for this lab calcium hydroxide is created then ionized Ca(OH)_2 when calcium is put in water reacts as follows $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$

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web k sp of calcium hydroxide 1 author j m mccormick last update ocober 13 2013 introduction an equilibrium constant k is related to ΔG for a process through eqn 1 and as long as ΔH and ΔS are independent of temperature eqn 1 can be combined with eqn 2 to determine ΔH and ΔS from the temperature dependence of k

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web sep 18 2020 in this experiment we will titrate a saturated calcium hydroxide solution with a standardized hcl solution we will determine the hydroxide ion concentration for the saturated calcium hydroxide

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web your primary objective in this experiment is to test a saturated solution of calcium hydroxide and use your observations and measurements to calculate the ksp of the compound you will do this by titrating the prepared ca oh 2 solution with a standard hydrochloric acid solution

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web in today s experiment you will determine the solubility product ksp of calcium hydroxide ca oh 2 by measuring the concentration of ca oh 2 in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of ca oh is represented in equation form as shown below ca oh s ca aq 2oh aq the solubility product expression describes in mathematical terms the equilibrium

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web ksp for calcium hydroxide calcium hydroxide is a soft white caustic powder used in making mortar cements calcium salts paints and petrochemicals it is also used in saltwater aquaria to make up kalkwasser limewater solutions for reef tanks and is used as a ph regulating agent notice that calcium hydroxide is divalent and thus has twice

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web use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the solution ca oh 2 à ca 2 2oh 5 x 10 2 m calcium ions 5 x 10 2 m 2 0 m hydroxide ions calculate calcium hydroxide ksp using the concentrations of calcium ions and hydroxide ions

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web madison bowe elana keller april 15 2014 introduction the objective of this lab was to measure the ksp of a slightly soluble basic salt and how solubility changes with temperature ksp is another form of equilibrium expression that deals with solubility the reaction that will take place is ca oh 2 ca2 2 oh

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web question experiment 8 determining ksp for calcium hydroxide post lab report sheet data table 2 points trial 1 trial 2 hci molarity 0 050 m volume of filtered ca oh 2 solution ml 5 00 5 00 3 71 13 75 equivalence point volume ml oh in saturated ca oh 2 m ca in saturated ca oh 2 m ksp of ca oh 2 average ksp of ca oh 2

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web the chemical equation used was $\text{Ca(OH)}_2 \text{ (s)} + \text{Ca}^{2+} \text{ (aq)} + 2\text{OH}^- \text{ (aq)}$ the solubility product was determined to be $K_{sp} \text{ Ca(OH)}_2$ from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

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web jul 12 2023 because the concentration of a pure solid such as $\text{Ca}_3\text{PO}_4\text{ (s)}$ is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore $K_{sp} \text{ Ca}_3\text{PO}_4\text{ (s)} = [\text{Ca}^{2+}]^3 [\text{PO}_4^{3-}]^2$

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web lab report 8 determination of calcium hydroxide the purpose of this was to titrate a saturated Ca(OH)_2 solution with a standard HCl solution and find the Ksp or the OH concentration was found and then the concentration was found and u for this experiment a titration of calcium hydroxide with hydrochloric acid

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web question determining the ksp of calciumhydroxide $\text{Ca(OH)}_2 \text{ (s)} + \text{Ca}^{2+} \text{ (aq)} + 2\text{OH}^- \text{ (aq)}$ for this experiment the titration of calciumhydroxide with hydrochloric acid was investigated determining the ksp of calcium hydroxide $\text{Ca(OH)}_2 \text{ (s)} + \text{Ca}^{2+} \text{ (aq)} + 2\text{OH}^- \text{ (aq)}$ for this experiment the titration of calcium

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web question chem 182 experiment 8 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water solution of Ca(OH)_2 is represented in equation form solution of calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous as shown below

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web therefore it was determined through error calculations and ease of use that the most reliable valid method to use to get the closest Ksp value to the theoretical Ksp for Ca was the determination through pH method 2 introduction calcium hydroxide is a strong base soluble in water

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classified as insoluble based on solubility rules are actually slightly soluble each of these insoluble compounds actually dissolves to

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we will filter the solid calcium hydroxide from the solution and then titrate the dissolved calcium hydroxide in solution with a known concentration of hydrochloric acid using titrimetric calculations we will determine the Ca^{2+} and OH^- concentrations in solution

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we calculate the ksp for calcium hydroxide $\text{ksp Ca}^{2+} \text{OH}^{-2} = 5 \times 10^{-6}$ find the accepted value of the ksp for calcium hydroxide and compare it with your value discuss the discrepancy and suggest possible sources of experimental error
accepted value for ksp for sodium hydroxide is 5×10^{-6}