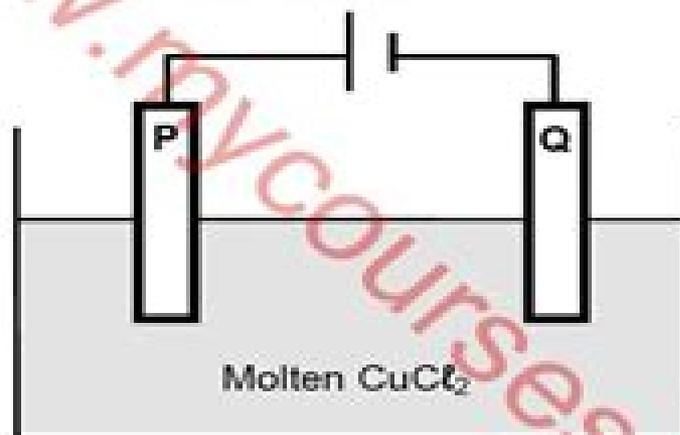


**QUESTION 5 (Start on a new page.)**

The diagram below represents the electrochemical cell used in the electrolysis of molten  $\text{CuCl}_2$ . **P** and **Q** are carbon electrodes.



- 5.1 Write down the magnitude of the copper charge in  $\text{CuCl}_2$ . (1)
- 5.2 Define the term *electrolysis*. (2)
- 5.3 Is the reaction ENDOTHERMIC or EXOTHERMIC? (1)
- 5.4 Which electrode is the cathode? Write down only **P** or **Q**. (1)
- 5.5 Write down the observations made at the following electrodes:
- 5.5.1 **P** (1)
- 5.5.2 **Q** (1)
- 5.6 Write down a balanced chemical equation for the net cell reaction of the above cell. (3)
- 5.7 Give a reason why  $\text{CuCl}_2$  is used in its molten form instead of its solid state. (2)

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**Rachel Sandford**



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