

**CHEMFILE MINI-GUIDE TO PROBLEM SOLVING**

*Mole Concept*

Suppose you want to carry out a reaction that requires combining one atom of iron with one atom of sulfur. How much iron should you use? How much sulfur? When you look around the lab, there is no device that can count numbers of atoms. Besides, the merest speck (0.001 g) of iron contains over a billion billion atoms. The same is true of sulfur.

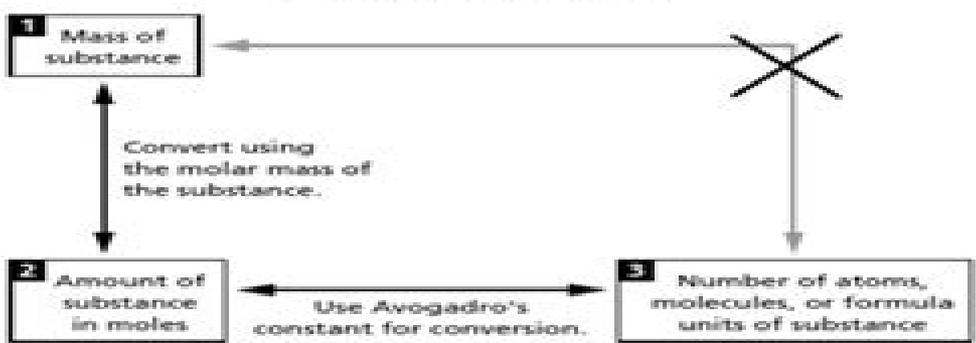
Fortunately, you do have a way to relate mass and numbers of atoms. One iron atom has a mass of 55.847 amu, and 55.847 g of iron contains  $6.022\ 137 \times 10^{23}$  atoms of iron. Likewise, 32.066 g of sulfur contains  $6.022\ 137 \times 10^{23}$  atoms of sulfur. Knowing this, you can measure out 55.847 g of iron and 32.066 g of sulfur and be pretty certain that you have the same number of atoms of each.

The number  $6.022\ 137 \times 10^{23}$  is called Avogadro's constant. For most purposes it is rounded off to  $6.022 \times 10^{23}$ . Because this is an awkward number to write over and over again, chemists refer to it as a *mole* (abbreviated mol).  $6.022 \times 10^{23}$  objects is called a mole, just as you call 12 objects a dozen.

Look again at how these quantities are related.

55.847 g of iron =  $6.022 \times 10^{23}$  iron atoms = 1 mol of iron  
 32.066 g of sulfur =  $6.022 \times 10^{23}$  sulfur atoms = 1 mol of sulfur

**General Plan for Converting Mass, Amount, and Numbers of Particles**



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# Holt Chemfile Mole Concept Answer Guide

**Ralph Thomas Myers**



### **Holt Chemfile Mole Concept Answer Guide:**

Holt Chemistry File ,1998 This reference is a must for students who need extra help reteaching or extra practice The guide moves students through the same concepts as the text but at a slower pace More descriptive detail along with visual algorithms provides a more structured approach Each chapter closes with a large bank of practice problems Book jacket

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