

Class 11

Chemistry Book

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STOICHIOMETRY Chapter

LEARNING OBJECTIVES

Students will be able to:

- Describe atomic and ionic radii, ionization enthalpy.
- Describe Avogadro's number and describe the relationship between mass and Avogadro number.
- Define molarity of ions, empirical formula and find its applications in writing chemical formulae.
- Perform stoichiometric calculations with limiting reagent using molar, empirical formula, percent and volumetric composition.
- Identify the limiting reagent in a reaction.
- Show the limiting reagent in a reaction and calculate the maximum amount of products produced and the amount of any excess reagent.
- Calculate theoretical yield, actual yield, percentage yield and atom economy.
- Calculate theoretical cell potential of a galvanic cell.

EXERCISES

The world is full of examples of chemical reactions. Can you imagine how useful it is to know the amount of substances formed when given known quantities of reactants? In a chemical change, there should be conservation of mass. Studying the quantities of various gases will be produced due to the combustion of gasoline in an internal combustion engine? It is important for you to know the amount of products formed in a reaction with adequate quantity of reactants and products. You need to calculate the amount of water produced in a reaction. You need to be able to compare the amount of water produced with the knowledge of stoichiometry.

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Michael Brown



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