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**Karin Nielsen-Saines**



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web now realizing that heating and cooling curves represent the amount of heat absorbed or released by a substance during phase changes remember heat uses the variable  $q$  here we have a heating curve versus a cooling curve in a heating curve our substance is absorbing heat

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web describe the processes represented by typical heating and cooling curves and compute heat flows and enthalpy changes accompanying these processes explain the construction and use of a typical phase diagram

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web heating and cooling curves a heating curve can be used to calculate the enthalpy when a substance is heated if we were to heat 25 00 g of water from 15 0 c to 115 0 c we can determine  $\Delta h$  for the heating process

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web determine the final temperature and state of the ammonia once the system reaches equilibrium assuming that no heat is

lost to the surroundings use the information below normal boiling point of  $\text{NH}_3$   $\Delta H_{\text{vap}} = 23.4 \text{ kJ mol}^{-1}$  specific heat  $\text{NH}_3$   $1.474 \text{ J g}^{-1} \text{ K}^{-1}$  specific heat  $\text{NH}_3$   $2.06 \text{ J g}^{-1} \text{ K}^{-1}$

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web apr 3 2023 the experiment described above can be summarized in a graph called a heating curve figure below figure 11 16 1 11 16 1 in the heating curve of water the temperature is shown as heat is continually added changes of state occur during plateaus because the temperature is constant

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web once all of the material has solidified at the time indicated by point c the cooling will continue at a rate determined by the heat capacities of the two solids a and b the composition and of course the geometry of the experimental set up by measuring cooling curves for samples of varying composition one can map the entire phase

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heating curves have five segments let s study each one by looking at heating ice to

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