

PART I - TRANSPORT PROCESSES: MOMENTUM, HEAT, AND MASS

(1.2-1)

Eq. (1.3-4) $K = 353.2 = ^\circ\text{C} + 273.15 \quad ^\circ\text{C} = 80.05 = \boxed{80^\circ\text{C}}$

Eq. (1.3-2) $^\circ\text{C} = 80.0 = (1/1.8)(^\circ\text{F} - 32) \quad ^\circ\text{F} = \boxed{176^\circ\text{F}}$

Eq. (1.3-3) $^\circ\text{R} = ^\circ\text{F} + 460 = 176 + 460 \quad ^\circ\text{R} = \boxed{636^\circ\text{R}}$

(1.2-2)

Eq. (1.3-2) $^\circ\text{C} = (1/1.8)(^\circ\text{F} - 32) = (1/1.8)(155 - 32) = \boxed{68.33^\circ\text{C}}$

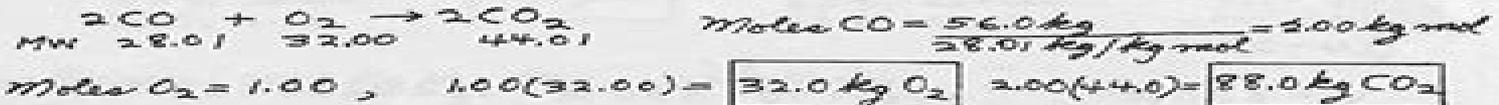
Eq. (1.3-4) $K = ^\circ\text{C} + 273.15 = 68.33 + 273.15 = \boxed{341.5\text{ K}}$

Eq. (1.3-3) $^\circ\text{R} = ^\circ\text{F} + 460 = 155 + 460 = \boxed{615^\circ\text{R}}$

(1.3-1)

$MW(\text{O}_2) = 32.00, \quad MW(\text{N}_2) = 28.02$
 $MW(\text{air}) = 0.21(32.00) + 0.79(28.02) = \boxed{28.9\text{ kg/kg mol}}$

(1.3-2)



(1.3-3)

Gas	g	MW	g mol	Mol free
N ₂	20	28.02	0.7138	0.1648 mol free 0.5990 0.2362 $\hline 1.0000$
O ₂	83	32.00	2.5938	
CO ₂	45	44.01	1.0225	
Total	148 g		4.3301 g mol	

$Average\ mol\ wt = \frac{148\text{ g}}{4.3301\text{ g mol}} = \boxed{34.2\text{ g/g mol}}$
 $\boxed{34.2\text{ kg/kg mol}}$

Geankoplis Transport Processes Solution 4th Manual

F. Carl Knopf



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Modeling, Analysis and Optimization of Process and Energy Systems F. Carl Knopf, 2011-12-14 Energy costs impact the profitability of virtually all industrial processes Stressing how plants use power and how that power is actually generated this book provides a clear and simple way to understand the energy usage in various processes as well as methods for optimizing these processes using practical hands on simulations and a unique approach that details solved problems utilizing actual plant data Invaluable information offers a complete energy saving approach essential for both the chemical and mechanical engineering curricula as well as for practicing engineers British Books in Print ,1985 **Solutions Manual to Accompany Transport Processes and Unit Operations, Second Edition, and Transport Processes** Christie J. Geankoplis, 1983 *Separation Process Engineering* Phillip C. Wankat, 2012 The Definitive Fully Updated Guide to Separation Process Engineering Now with a Thorough Introduction to Mass Transfer Analysis Separation Process Engineering Third Edition is the most comprehensive accessible guide available on modern separation processes and the fundamentals of mass transfer Phillip C Wankat teaches each key concept through detailed realistic examples using real data including up to date simulation practice and new spreadsheet based exercises Wankat thoroughly covers each of today s leading approaches including flash column and batch distillation exact calculations and shortcut methods for multicomponent distillation staged and packed column design absorption stripping and more In this edition he also presents the latest design methods for liquid liquid extraction This edition contains the most detailed coverage available of membrane separations and of sorption separations adsorption chromatography and ion exchange Updated with new techniques and references throughout Separation Process Engineering Third Edition also contains more than 300 new homework problems each tested in the author s Purdue University classes Coverage includes Modular up to date process simulation examples and homework problems based on Aspen Plus and easily adaptable to any simulator Extensive new coverage of mass transfer and diffusion including both Fickian and Maxwell Stefan approaches Detailed discussions of liquid liquid extraction including McCabe Thiele triangle and computer simulation analyses mixer settler design Karr columns and related mass transfer analyses Thorough introductions to adsorption chromatography and ion exchange designed to prepare students for advanced work in these areas Complete coverage of membrane separations including gas permeation reverse osmosis ultrafiltration pervaporation and key applications A full chapter on economics and energy conservation in distillation Excel spreadsheets offering additional practice with problems in distillation diffusion mass transfer and membrane separation **The British Library General Catalogue of Printed Books 1976 to 1982** British Library, 1983 **Scientific and Technical Books and Serials in Print** ,1984 *Books in Print Supplement* ,1982 *Chemical Engineering Progress* ,1994 **Journal of Food Protection** ,1992 Choice Richard K. Gardner, Phyllis Grumm, 1976 Subject Guide to Books in Print ,1984

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Dahlan, 2015 On air pollution and environmental protection in Malaysia *Transport Processes and Separatn*

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buffers typically consist of an acid base pair with the acid and base differing by the presence or absence of a proton a conjugate acid base pair for instance one of the buffers that maintain the ph of human blood involves carbonic acid H_2CO_3 and its conjugate base the bicarbonate ion HCO_3^-

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acids bases and enzymes many acids and bases in living things provide the ph that enzymes need enzymes are biological catalysts that must work effectively for biochemical reactions to occur most enzymes can do their job only at a certain level of acidity cells secrete acids and bases to maintain the proper ph for enzymes to do their work

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acid base titration difference between acid and base acids and bases definition an acid is any hydrogen containing substance that is capable of donating a proton hydrogen ion to another substance a base is a molecule or ion able to accept a hydrogen ion from an acid acidic substances are usually identified by their sour taste

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a binary acid is an acid that consists of hydrogen and one other element the most common binary acids contain a halogen the acid name begins with the prefix hydro followed by the base name of the anion followed by the suffix ic

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oct 18 2023 most of the acids are solid it has a sour taste they produce hydrogen after reacting with active metals like zinc magnesium aluminium or iron acids have a low ph value i e below 7 salts and water are produced when an acid reacts with base or alkalis most of the acids are solids

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arrhenius s definition of acids and bases the earliest definition of acids and bases is arrhenius s definition which states that an acid is a substance that forms hydrogen ions h when dissolved in water and a base is a substance that forms hydroxide ions oh when dissolved in water

10 1 acids properties and examples chemistry libretxts - May 20 2023

mar 24 2021 acids react with bases to produce a salt compound and water when equal moles of an acid and a base are combined the acid is neutralized by the base the products of this reaction are an ionic compound which is labeled as a salt and water

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the reaction of an acid and a base is called a neutralization reaction although acids and bases have their own unique chemistries the acid and base cancel each other s chemistry to produce a rather innocuous substance water in fact the general reaction between an acid and a base is ce acid base water salt nonnumber

14 2 acids properties and examples chemistry libretxts - Apr 19 2023

may 20 2018 acids react with bases to produce a salt compound and water when equal moles of an acid and a base are combined the acid is neutralized by the base the products of this reaction are an ionic compound which is labeled as a salt and water

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nov 13 2022 suggest simple tests you could carry out to determine if an unknown substance is an acid or a base state the chemical definitions of an acid and a base in terms of their behavior in water write the formula of the salt formed when a given acid and base are combined