

General Chemistry II Jasperse
Acid-Base Chemistry: Extra Practice Problems

General Types/Groups of problems:

Conceptual Questions: Acids, Bases, and Conjugates, Miscellaneous	p1	K_b and pK_b , Base Strength, and using K_b or pK_b to Calculate $[OH^-]$, pOH , pH , and/or $[H^+]$	p7-10
Recognizing Strong versus Weak Acids; Recognizing Basic versus Nonbasic	p3	Recognizing Acid/Base Properties when Ionics are Dissolved in Water	p11
pH Calculations: Relationships between pH and pOH	p4	Answers	p12
K_a , Sense + Calculations: Using K_a or pK_a to Calculate $[H^+]$ and/or pH; using pH to calculate K_a or pK_a	p5		

Conceptual Questions: Acids, Bases, and Conjugates, Miscellaneous

- In the Brønsted-Lowry definition of acids and bases, an acid _____
 - is a proton donor.
 - is a proton acceptor.
 - forms stable hydrogen bonds.
 - breaks stable hydrogen bonds.
 - corrodes metals.
- In the Brønsted-Lowry definition of acids and bases, a base _____
 - is a proton donor.
 - is a proton acceptor.
 - forms stable hydrogen bonds.
 - breaks stable hydrogen bonds.
 - corrodes metals.
- In the following reaction in aqueous solution, the acid reactant is _____ and its conjugate base product is _____

$$CH_3COOH + NH_3 \rightleftharpoons CH_3COO^- + NH_4^+$$
 - CH_3COOH ; CH_3COO^-
 - CH_3COOH ; NH_4^+
 - NH_3 ; CH_3COO^-
 - NH_3 ; NH_4^+
 - CH_3COOH ; H_3O^+
- In the following reaction in aqueous solution, the acid reactant is _____, and its conjugate base product is _____

$$CH_3NH_2 + HSO_4^- \rightleftharpoons CH_3NH_3^+ + SO_4^{2-}$$
 - CH_3NH_2 ; $CH_3NH_3^+$
 - CH_3NH_2 ; SO_4^{2-}
 - HSO_4^- ; $CH_3NH_3^+$
 - HSO_4^- ; SO_4^{2-}
 - HSO_4^- ; H_3O^+
- Which of the following is the conjugate acid of the hydrogen phosphate ion, HPO_4^{2-} ?
 - $H_2PO_4^-$
 - H_3PO_4
 - HPO_4^{2-}
 - PO_4^{3-}
 - H_2O^+

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Princeton Review (Firm)



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