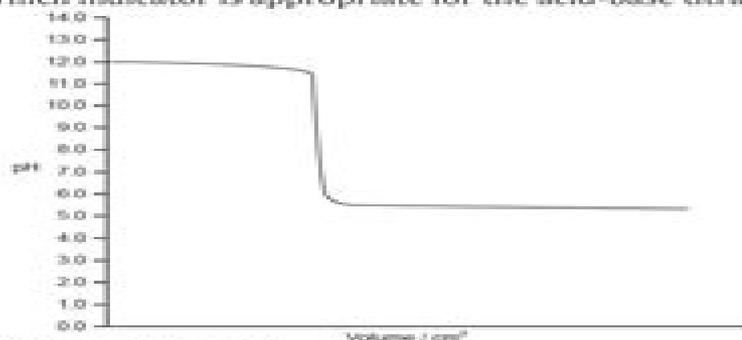


8. Which indicator is appropriate for the acid-base titration shown below?



- A. Thymol blue ($pK_a = 1.5$)
- B. Methyl orange ($pK_a = 3.7$)
- C. Bromophenol blue ($pK_a = 4.2$)
- D. Phenolphthalein ($pK_a = 9.6$)

9. Which species produced by the successive dissociations of phosphoric acid, H_3PO_4 , are amphiprotic?

- A. HPO_4^{2-} and PO_4^{3-}
- B. $H_2PO_4^-$ and HPO_4^{2-}
- C. $H_2PO_4^-$ and PO_4^{3-}
- D. HPO_4^{2-} only

10. What is the pH of 1.0×10^{-3} mol dm^{-3} sodium hydroxide, $NaOH(aq)$? $K_w = 1.0 \times 10^{-14}$

- A. 3
- B. 4
- C. 10
- D. 11

11. Which species acts as a Lewis and Brønsted-Lowry base?

- A. $[Al(H_2O)_6]^{3+}$
- B. BF_3
- C. NH_4^+
- D. OH^-

12. A buffer is produced by mixing 20.0 cm^3 of 0.10 mol dm^{-3} ethanoic acid, $CH_3COOH(aq)$, with 0.10 mol dm^{-3} sodium hydroxide, $NaOH(aq)$. What is the volume of $NaOH$ required and the pH of the buffer?

	Volume of $NaOH / \text{cm}^3$	pH of buffer
A.	40.0	9.2
B.	40.0	4.8
C.	10.0	9.2
D.	10.0	4.8

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L Darling-Hammond



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