

Final Exam Study Guide (General Chemistry I)

- Chemistry - the study of matter and the changes that matter undergoes
- Matter - anything that has mass and takes up space
 - Three states of matter: solid, liquid, and gas (plasma is a gas)

	Proximity of Particles	Energetics	Size/Shape
Solid	Closely packed and in direct contact with each other	Have very little energy of motion (slow)	Fixed shape and fixed volume
Liquid	In indirect contact with each other; usually about a particles distance apart (bounce off each other)	Have an intermediate amount energy of motion (medium)	Does not have a fixed shape; has a fixed volume
Gas	Particles are very far apart	Have a large amount of energy of motion (fast)	No fixed shape or volume

- **PARTICLES ARE ALWAYS MOVING!**
- "Can you undo it?" (Page 12)
 - Physical Changes - changes physical appearance but not its composition
 - Ex. Melting, evaporation, freezing, etc.
 - Chemical Changes - changes in chemical composition of matter (chemical reaction)
 - Ex. Baking soda and vinegar, burning, mixing, etc.
 - Physical Properties - characteristics that can be described by observing the matter; no change in composition is required to detect a physical property
 - Ex. Color, odor, density, etc.
 - Chemical Properties - a change in composition is required to detect a chemical property
 - Ex. Flammability
- Accuracy Versus Precision:
 - In every measurement, there is some degree of uncertainty.
 - **NUMBERS WITHOUT UNITS ARE MEANINGLESS IN MEASUREMENT!**
 - Accuracy - the agreement of a measurement with a known or accepted value (correctness)
 - Precision - the agreement of a measurement with other measurements in a given set of measurements (reproducibility/repeatability)
 - You can have accuracy without precision, and precision without accuracy.
 - **WE ARE ONLY AS CERTAIN AS OUR LEAST CERTAIN NUMBER!**
 - 254 can mean 253.5 to 254.4
 - 254.0 can mean 253.95 to 254.04
 - 254.00 can mean 253.950 to 254.004
 - 254.000 can mean 253.9500 to 254.0004
- Density - the mathematical relationship of mass to volume

General Chemistry 1 Final Exam Study Guide

Aftab S. Hassan



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Directory of Distance Learning Opportunities Modoc Press, Inc.,2003-02-28 This book provides an overview of current K 12 courses and programs offered in the United States as correspondence study or via such electronic delivery systems as satellite cable or the Internet The Directory includes over 6 000 courses offered by 154 institutions or distance learning consortium members Following an introduction that describes existing practices and delivery methods the Directory offers three indexes Subject Index of Courses Offered by Level Course Level Index Geographic Index All information was supplied by the institutions Entries include current contact information a description of the institution and the courses offered grade level and admission information tuition and fee information enrollment periods delivery information equipment requirements credit and grading information library services and accreditation

Instructor's manual and solutions to problems to accompany General chemistry: principles and structures James E. Brady,1975-04 Journal of Developmental Education ,2015

El-Hi Textbooks and Serials in Print, 1996 R R Bowker Publishing,Bowker,1996-02 New York Herald Tribune Book Review ,1959

Medical Books and Serials in Print ,1983 *Selections from Communication Teacher, (2004, Printed) to Accompany the Art of Public Speaking* Jr. Lucas, Jr.,2003-06 Cornell University Courses of Study Cornell University,1992

Subject Guide to Books in Print ,2001 *El-Hi Textbooks in Print* ,1984

DAT Aftab S. Hassan,2000 This study guide for the new content and skills tested on the Dental Admission Test DAT offers an extensive perpetual ability workout in one two and three dimensions Extensive practice items reflect the DAT model and a full length practice test is included

Chemistry George M. Bodner,Harry L. Pardue,1989-01-23 A modern experimental approach to first year chemistry This unique introductory account employs experimental observations to construct the principles of general chemistry An early introduction to observable descriptive chemistry lays the basis for the well developed exposition that follows

Materials Performance ,1999-07 *Guide to American Graduate Schools* ,1982

The 2002 Guide to the Evaluation of Educational Experiences in the Armed Services American Council on Education,2002 Long considered to be the standard reference work in this area this three volume set describes more than 8 000 courses offered between January 1990 and the present by various service branches and the Department of Defense Long considered to be the standard reference work in this area this three volume set describes more than 8 000 courses offered between January 1990 and the present by various service branches and the Department of Defense Updated every two years

The 2004 Guide to the Evaluation of Educational Experiences in the Armed Services American Council on Education Staff,2004-10-27

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Geology Edward C. Gruber,1963

Basic Chemistry Steven S. Zumdahl,1999 For a full description see catalog entry for Zumdahl Introductory Chemistry A Foundation

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In an era characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its ability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**General Chemistry 1 Final Exam Study Guide**," a mesmerizing literary creation penned by a celebrated wordsmith, readers attempt an enlightening odyssey, unraveling the intricate significance of language and its enduring affect our lives. In this appraisal, we shall explore the book is central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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