

AP Chemistry: Equilibrium Multiple Choice

57. For the reaction $A_{(g)} \rightleftharpoons B_{(g)} + C_{(g)}$, the equilibrium constant, K_p , is 2×10^{-4} at 25°C . A mixture of the three gases at 25°C is placed in a reaction flask and the initial pressures are $P_A = 2 \text{ atm}$, $P_B = 0.5 \text{ atm}$, and $P_C = 1 \text{ atm}$. At the instant of mixing, which of the following is true for the reaction as written?

- (A) $\Delta G < 0$ (B) $\Delta G > 0$ (C) $\Delta S = 0$ (D) $\Delta G^\circ = 0$ (E) $\Delta G^\circ < 0$



Consider the equilibrium above. Which of the following changes will increase the concentration of HgI_4^{2-} ?

- (A) Increasing the concentration of OH^- (B) Adding 6 M HNO_3
(C) Increasing the mass of HgO present (D) Increasing the temperature (E) Adding a catalyst

29. In which of the following systems would the number of moles of the substances present at equilibrium NOT be shifted by a change in the volume of the system at constant temperature?

- (A) $\text{CO}_{(g)} + \text{NO}_{(g)} \rightleftharpoons \text{CO}_{2(g)} + 1/2 \text{N}_{2(g)}$ (B) $\text{N}_{2(g)} + 3 \text{H}_{2(g)} \rightleftharpoons 2 \text{NH}_{3(g)}$
(C) $\text{N}_{2(g)} + 2 \text{O}_{2(g)} \rightleftharpoons 2 \text{NO}_{2(g)}$ (D) $\text{N}_2\text{O}_{4(g)} \rightleftharpoons 2 \text{NO}_{2(g)}$
(E) $\text{NO}_{(g)} + \text{O}_{3(g)} \rightleftharpoons \text{NO}_{2(g)} + \text{O}_{2(g)}$



Some PCl_3 and Cl_2 are mixed in a container at 200°C and the system reaches equilibrium according to the equation above. Which of the following causes an increase in the number of moles of PCl_3 present at equilibrium?

- I. Decreasing the volume of the container
II. Raising the temperature
III. Adding a mole of He gas at constant volume
(A) I only (B) II only (C) I and III only (D) II and III only (E) I, II, and III



Equal numbers of moles of HCl and O_2 in a closed system are allowed to reach equilibrium as represented by the equation above. Which of the following must be true at equilibrium?

- I. $[\text{HCl}]$ must be less than $[\text{Cl}_2]$.
II. $[\text{O}_2]$ must be greater than $[\text{HCl}]$.
III. $[\text{Cl}_2]$ must equal $[\text{H}_2\text{O}]$.
(A) I only (B) II only (C) I and III only (D) II and III only (E) I, II, and III

65. The solubility of CuI is 2×10^{-8} molar. What is the solubility product constant, K_{sp} , for CuI ?

- (A) 1.4×10^{-3} (B) 2×10^{-8} (C) 4×10^{-12} (D) 2×10^{-12} (E) 8×10^{-18}

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