

## General Chemistry II Equation Sheet

Think about how to set up the problem first, then apply the needed principles and formulas.

### Phase Changes

$$\Delta T_b = iK_b m$$

$$\Delta T_f = iK_f m$$

$$\pi = iMRT$$

$$c = kP$$

$$P_{\text{solution}} = X_{\text{solvent}} P_{\text{solvent}}^{\circ}$$

$$\Delta P = P_{\text{solvent}}^{\circ} - P_{\text{solution}} = X_{\text{solute}} P_{\text{solvent}}^{\circ}$$

### Cells

$$E = E^{\circ} - \frac{RT}{nF} \ln(Q)$$

$$E = E^{\circ} - \frac{0.0592}{n} \log(Q) \text{ (at } 25^{\circ}\text{C)}$$

$$\Delta G = -nFE_{\text{cell}}$$

$$\Delta G^{\circ} = -nFE_{\text{cell}}^{\circ}$$

$$E_{\text{cell}}^{\circ} = \frac{RT}{nF} \ln(K)$$

$$E_{\text{cell}}^{\circ} = \frac{0.0592V}{n} \log(K) \text{ (at } 25^{\circ}\text{C)}$$

USE ONE:

$$E_{\text{cell}}^{\circ} = E_{\text{cathode}}^{\circ}(\text{red.}) - E_{\text{anode}}^{\circ}(\text{red.})$$

$$E_{\text{cell}}^{\circ} = E_{\text{cathode}}^{\circ}(\text{red.}) + E_{\text{anode}}^{\circ}(\text{red.})$$

### Thermodynamics

$$\Delta G = \Delta H - T\Delta S$$

$$\Delta S_{\text{surroundings}} = \frac{-\Delta H_{\text{sys}}}{T}$$

$$\Delta S_{\text{universe}} = \Delta S_{\text{sys}} + \Delta S_{\text{surroundings}}$$

$$\Delta S_{\text{rxn}}^{\circ} = \sum S_{\text{products}}^{\circ} - \sum S_{\text{reactants}}^{\circ}$$

$$\Delta G_{\text{rxn}}^{\circ} = \sum \Delta G_{\text{products}}^{\circ} - \sum \Delta G_{\text{reactants}}^{\circ}$$

### Concentrations

$$\text{molality} = \frac{\text{moles}_{\text{solute}}}{\text{mass}_{\text{solvent}}(\text{kg})}$$

$$\% \text{ by mass} = \frac{\text{moles}_{\text{solute}} \times \text{molar mass}_{\text{solute}}}{\text{moles}_{\text{solute}} \times \text{molar mass}_{\text{solute}} + \text{mass}_{\text{solvent}}} \times 100\%$$

### Acids and Bases

$$K_w = [H_3O^+][OH^-] = K_a \times K_b$$

$$pOH = -\log[OH^-]$$

$$[OH^-] = 10^{-pOH}$$

$$pH + pOH = pK_a + pK_b = 14.00 \text{ (at } 25^{\circ}\text{C)}$$

$$pH = pK_a + \log \left[ \frac{[A^-]}{[HA]} \right]$$

$$pOH = pK_b + \log \left[ \frac{[OH^-]}{[B]} \right]$$

$$\% \text{ ionization} = \frac{[H_3O^+]_{\text{eq}}}{[HA]_0} \times 100\%$$

$$pK_a = -\log(K_a)$$

$$pK_b = -\log(K_b)$$

$$K_a = 10^{-pK_a}$$

$$K_b = 10^{-pK_b}$$

### Equilibria

$$Q_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

$$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

$$Q_p = \frac{P_C^c P_D^d}{P_A^a P_B^b}$$

$$K_p = \frac{P_C^c P_D^d}{P_A^a P_B^b}$$

$$K_p = K_c [R \times T]^{\Delta n}$$

$$\Delta G = \Delta G^{\circ} + RT \ln(Q)$$

$$\Delta G^{\circ} = -RT \ln(K)$$

$$\text{rate} = k[A]^m[B]^n$$

$$k = Ae^{-\frac{E_a}{RT}}$$

$$\ln\left(\frac{t_2}{t_1}\right) = \frac{E_a}{R} \left(\frac{1}{T_1} - \frac{1}{T_2}\right)$$

$$\text{half life} = \frac{\ln(2)}{k}$$

$$[A]_t = -kt + [A]_0$$

$$\ln([A]_t) = -kt + \ln([A]_0)$$

$$\frac{1}{[A]_t} = kt + \frac{1}{[A]_0}$$

# General Chemistry Ii Cheat Guide

**JS Bruner**



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