

SS2 CHEMISTRY FIRST TERM SCHEME OF WORK

1. PERIODIC TABLE [GROUPS AND PERIODS], PROPERTIES
2. PERIODIC TABLE [FAMILIES OF ELEMENTS]
3. MASS AND VOLUME RELATIONSHIPS IN REACTIONS [ACIDS AND BASES TITRATION PRACTICAL]
4. MASS AND VOLUME RELATIONSHIPS IN REACTIONS
5. CONTINUOUS ASSESSMENT
6. VOLUMETRIC AND QUALITATIVE ANALYSIS – PRACTICAL
7. MID-TERM BREAK
8. TYPES OF CHEMICAL REACTIONS
9. OXIDATION AND REDUCTION
10. OXIDATION AND REDUCTION – REDOX REACTION
11. ELECTROLYSIS
12. REVISION
13. EXAMINATION.

THE PERIODIC TABLE

Periodic means recur at regular intervals. When elements are systematically arranged in order of ascending atomic number, certain characteristics recur at regular intervals.

Dmitri Ivanovich Mendeleev, a Russian chemist, published the first version of the periodic table in 1869. With this arrangement, certain elements with similar chemical properties and electronic structures appear in a given group.

The periodic law states that the properties of the elements are the periodic function of their atomic number.

Periodic table

The elements that appear in a vertical column are said to belong to the same group or family. They have the same number of outer electrons or valence electrons and therefore exhibit similar chemical behaviour. The elements in the same horizontal row are said to belong to the same period.

Period- Elements in the same row are in the same period and have the same number of shells.

The elements in group I are Lithium (Li), sodium (Na), potassium (K), rubidium (Rb), Caesium (Cs) and Francium (Fr). These are known as the alkali metals.

The elements in group II are called the alkaline earth metals. The alkaline earth metals are beryllium

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M Woodhall



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