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 $\text{GaBr} + \text{NaBr} + \text{SnO}_4 + \text{Na}_2\text{SO}_3 + \text{Ga}_2\text{SO}_3 + 3\text{SO}_3 + 3\text{H}_2\text{O} + 3\text{NF}_3 + 3\text{SnF}_2 + \text{N}_2\text{O}_3$ using the equation from problem 2 above answer the following

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 $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$
 $\text{C}_6\text{H}_{10}\text{O}_2 + 17\text{O}_2 \rightarrow 12\text{CO}_2 + 10\text{H}_2\text{O}$
 $3\text{HBr} + \text{KHCO}_3 + \text{H}_2\text{O} + \text{KBr} + \text{CO}_2 \rightarrow 2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 + \text{Ga}_2\text{SO}_3 + 3\text{SO}_3 + 6\text{NaBr} + 3\text{SnO}_2 + 3\text{NF}_3 + 3\text{SnF}_2 + \text{N}_2\text{O}_3$ using the following equation
 $2\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Na}_2\text{SO}_4$

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 $\text{Ce} + \text{Zn} + \text{CuCl}_2 \rightarrow \text{ZnCl}_2 + \text{Cu} + \text{Zn} + \text{CuCl}_x$
 $2\text{ZnCl}_x + 2\text{Cu}$ how many moles of $\text{Ce} + \text{ZnCl}_2 + \text{ZnCl}_x + 2$ will be produced from 23.0 g of $\text{Ce} + \text{Zn}$ assuming $\text{Ce} + \text{CuCl}_2 + \text{CuCl}_x + 2$ is available in excess

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 $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$
 $\text{C}_6\text{H}_{10}\text{O}_2 + 17\text{O}_2 \rightarrow 12\text{CO}_2 + 10\text{H}_2\text{O}$
 $3\text{HBr} + \text{KHCO}_3 + \text{H}_2\text{O} + \text{KBr} + \text{CO}_2 \rightarrow 2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 + \text{Ga}_2\text{SO}_3 + 3\text{SO}_3 + 6\text{NaBr}$
 $5\text{SnO}_2 + 3\text{NF}_3 + 3\text{SnF}_2 + \text{N}_2\text{O}_3$ using the following equation
 $2\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Na}_2\text{SO}_4$

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