

Atoms, Bonds and Groups

Atoms, Isotopes and Relative Atomic Masses

70 Marks

1. Isotopes of europium have differences and similarities.

(i) In terms of protons, neutrons and electrons, how is an atom of ^{151}Eu **different** from an atom of ^{153}Eu ?

.....
.....

[1]

(ii) In terms of protons, neutrons and electrons, how is an atom of ^{151}Eu **similar** to an atom of ^{153}Eu ?

.....
.....

[1]

[Total 2 marks]

2. Europium, atomic number 63, is used in some television screens to highlight colours. A chemist analysed a sample of europium using mass spectrometry. The results are shown in the table below.

isotope	relative isotopic mass	abundance (%)
^{151}Eu	151.0	47.77
^{153}Eu	153.0	52.23

(a) Define the term *relative isotopic mass*.

.....
.....
.....
.....

[2]

Friday 23rd May 2014 Atoms Bonds And Groups

Alan Holden



Friday 23rd May 2014 Atoms Bonds And Groups:

The Annalist, 1932 *Bonds Between Atoms* Alan Holden, 1971 Atoms, Chemical Bonds, and Bond Dissociation Energies Sándor Fliszár, 1994

Chemical bonds their intrinsic energies in ground state molecules and the energies required for their actual cleavage are the subject of this book. The theory modelled after a description of valence electrons in isolated atoms explains how intrinsic bond energies depend on the amount of electronic charge carried by the bond forming atoms. It also explains how bond dissociation depends on these charges. While this theory vividly explains thermochemical stability, future research could benefit from a better understanding of bond dissociation if we learn how the environment of a molecule affects its charges; we also learn how it modifies bond dissociation in that molecule. This essay is aimed at theoretical and physical organic chemists who are looking for new perspectives to old problems.

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