

Consider the data in the table.

Compound	Melting point (°C)	ΔH_{fus} (kJ/mol)	Boiling point (°C)	ΔH_{vap} (kJ/mol)
HF	-83.11	4.577	19.54	25.18
HCl	-114.3	1.991	-84.9	17.53
HBr	-86.96	2.406	-67.0	19.27
HI	-50.91	2.871	-35.38	21.16

Using the data in the table, calculate ΔS_{fus} and ΔS_{vap} for HCl.

$\Delta S_{\text{fus}} =$ J/(K · mol)

$\Delta S_{\text{vap}} =$ J/(K · mol)

Determine the entropy change when 6.10 mol HCl(g) condenses at atmospheric pressure.

$\Delta S =$ J/K

General Chemistry Mcquarrie Answers

Clemens Wendtner



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