

## EQUILIBRIUM CONSTANT - WORKSHEET 2\_ANSWERS

1. Write the equilibrium constant expression for each of the following reactions:



$$K = \frac{[NO_2]^2}{[N_2][O_2]}$$



$$K = \frac{[C_2H_6]}{[H_2]^3}$$



$$K = \frac{1}{[I_2(aq)]}$$

2. For the following reaction:



At equilibrium,  $[Fe^{2+}] = 0.50 M$ ,  $[Ag^+] = 1.0 M$ ,  $[Fe^{3+}] = 1.50 M$

a) Calculate the equilibrium constant for the reaction.

$$K = 3.0$$

b) If at equilibrium,  $[Fe^{2+}] = 0.20 M$  and  $[Ag^+] = 0.30 M$ , what must be the concentration of  $Fe^{3+}$ ?

$$[Fe^{3+}] = 0.18 \text{ mol/L}$$



Calculate the equilibrium concentration of oxygen gas if the NO and  $NO_2$  concentrations are equal at equilibrium.

$$[O_2] = 2.0 \times 10^{-6} \text{ mol/L}$$



Which of the above reactions would have the greatest amount of products present at equilibrium compared to reactants?

Reaction b. Since K is larger the amount of products must be larger, because K is products/reactants.

5. For the reaction:



at an equilibrium temperature of  $1000^\circ\text{C}$ , a 1.00 L flask contains 0.120 moles of ammonia, 1.03 moles of nitrogen and 1.62 moles of hydrogen. Calculate the equilibrium constant at this temperature.

$$K = 3.29 \times 10^{-2}$$

# Genetic Equilibrium Worksheet Answer Key

**IM Harris**



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