

Chapter Three: Basic Principles of Chemistry

EXERCISES AND PROBLEMS

Problem 3.1

1. What was Mendeleev's approach to the periodicity of chemical properties?
Mendeleev rearranged the known elements. He placed each new element (those not yet discovered) in a column (group) vertically opposite an element whose properties were most similar to those of the element. He placed the elements in order of increasing atomic weight horizontally. He placed the elements in order of increasing atomic weight vertically. He placed the elements in order of increasing atomic weight horizontally. He placed the elements in order of increasing atomic weight vertically. He placed the elements in order of increasing atomic weight horizontally. He placed the elements in order of increasing atomic weight vertically.
2. What is the periodic law? How was periodicity used to predict the properties of elements?
The periodic law states that the properties of the elements are a periodic function of their atomic number. Mendeleev used the periodic law to predict the properties of elements that had not yet been discovered.

Problem 3.2

1. What is the periodic law? How was periodicity used to predict the properties of elements?
The periodic law states that the properties of the elements are a periodic function of their atomic number. Mendeleev used the periodic law to predict the properties of elements that had not yet been discovered.
2. How was Mendeleev's periodic table used to predict the properties of elements?
Mendeleev's periodic table was used to predict the properties of elements that had not yet been discovered. He placed the elements in order of increasing atomic weight horizontally. He placed the elements in order of increasing atomic weight vertically. He placed the elements in order of increasing atomic weight horizontally. He placed the elements in order of increasing atomic weight vertically.
3. What is the periodic law? How was periodicity used to predict the properties of elements?
The periodic law states that the properties of the elements are a periodic function of their atomic number. Mendeleev used the periodic law to predict the properties of elements that had not yet been discovered.

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Albert A Gayle



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