

Quiz & Worksheet - Limiting Reactants & Excess Reactants

1. What is a limiting reactant?

- The reactant that is used up first and prevents more product from being made.
- The reactant that is used up last and prevents more product from being made.
- The reactant that makes the product.
- The substance that is in excess that doesn't get used up as a reactant.

2. Say you take a reactant A and calculate the amount of moles of another reactant B required to use up all of A. How do you know which of two reactants is the limiting one?

- You compare the calculated amount of B to the actual amount available. If more is required, then B is the limiting reactant. If less is required, then A is the limiting reactant.
- You compare the calculated amount of B to the actual amount available. If more is required, then A is the limiting reactant. If less is required, then B is the limiting reactant.
- If the calculated value of B is larger than the amount of A, then A is the limiting reactant.
- If the calculated value of B is larger than the amount of A, then B is the limiting reactant.

3. If 2.5 moles of zinc react with 6.0 moles of hydrochloric acid in the equation $\text{Zn} + 2\text{HCl} = \text{ZnCl}_2 + \text{H}_2$, what is the limiting reactant?

- Zn
- HCl
- ZnCl
- ZnH

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This lesson plan covers analyzing chemical reactions in order to determine limiting reactants and excess reactants including calculating the amount of excess reactant calculate the theoretical yield of a reaction and calculating the percent yield of a reaction

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