

## **Limiting Reagent Worksheet #1**

1. Given the following reaction: (Balance the equation first!)



- If you start with 14.8 g of  $\text{C}_3\text{H}_8$  and 3.44 g of  $\text{O}_2$ , determine the limiting reagent
- determine the number of moles of carbon dioxide produced
- determine the number of grams of  $\text{H}_2\text{O}$  produced
- determine the number of grams of excess reagent left

2. Given the following equation:



- If 10.0 g of  $\text{Al}_2(\text{SO}_4)_3$  is reacted with 10.0 g of  $\text{NaOH}$ , determine the limiting reagent
- Determine the number of moles of  $\text{Al}(\text{OH})_3$  produced
- Determine the number of grams of  $\text{Na}_2\text{SO}_4$  produced
- Determine the number of grams of excess reagent left over in the reaction

3. Given the following equation:



- If 25.4 g of  $\text{Al}_2\text{O}_3$  is reacted with 10.2 g of  $\text{Fe}$ , determine the limiting reagent
- Determine the number of moles of  $\text{Al}$  produced
- Determine the number of grams of  $\text{Fe}_3\text{O}_4$  produced
- Determine the number of grams of excess reagent left over in the reaction

# Limiting Reagent Question And Answer Key

**Edward I. Peters**



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