

Limiting Reactant Worksheet #2

1. Given the following reaction: (Balance the equation first!)



- If you start with 14.8 g of C_3H_8 and 3.44 g of O_2 , determine the limiting reactant.
- Determine the number of moles of carbon dioxide produced.
- Determine the number of grams of H_2O produced.
- Determine the number of grams of excess reactant left.

2. Given the following equation:



- If 10.0 g of $\text{Al}_2(\text{SO}_4)_3$ is reacted with 10.0 g of NaOH , determine the limiting reactant.
- Determine the number of moles of $\text{Al}(\text{OH})_3$ produced.
- Determine the number of grams of Na_2SO_4 produced.
- Determine the number of grams of excess reactant left over in the reaction.

3. Given the following equation:



- If 25.4 g of Al_2O_3 is reacted with 10.2 g of Fe , determine the limiting reactant.
- Determine the number of moles of Al produced.
- Determine the number of grams of Fe_3O_4 produced.
- Determine the number of grams of excess reactant left over in the reaction.

4. Consider the reaction



- 80.0 grams of I_2O_5 , reacts with 28.0 grams of carbon monoxide, CO . Determine the mass of iodine I_2 , which could be produced?
- If, in the above situation, only 0.160 moles, of iodine, I_2 was produced.
 - What mass of iodine was produced?
 - What percentage yield of iodine was produced?

5. Zinc and sulfur react to form zinc sulfide according to the equation



If 25.0 g of zinc and 30.0 g of sulfur are mixed,

- Which chemical is the limiting reactant?
- How many grams of ZnS will be formed?
- How many grams of the excess reactant will remain after the reaction is over?

6. Which element is in excess when 3.00 grams of Mg is ignited in 2.20 grams of pure oxygen? What mass is in excess? What mass of MgO is formed?

7. How many grams of Al_2S_3 are formed when 5.00 grams of Al is heated with 10.0 grams S ?

Limiting Reagent Worksheet 202 Answer Key

Christian G. Meyer



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Table of Contents Limiting Reagent Worksheet 202 Answer Key

1. Understanding the eBook Limiting Reagent Worksheet 202 Answer Key
 - The Rise of Digital Reading Limiting Reagent Worksheet 202 Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reagent Worksheet 202 Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limiting Reagent Worksheet 202 Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reagent Worksheet 202 Answer Key
 - Personalized Recommendations
 - Limiting Reagent Worksheet 202 Answer Key User Reviews and Ratings
 - Limiting Reagent Worksheet 202 Answer Key and Bestseller Lists
5. Accessing Limiting Reagent Worksheet 202 Answer Key Free and Paid eBooks
 - Limiting Reagent Worksheet 202 Answer Key Public Domain eBooks
 - Limiting Reagent Worksheet 202 Answer Key eBook Subscription Services
 - Limiting Reagent Worksheet 202 Answer Key Budget-Friendly Options
6. Navigating Limiting Reagent Worksheet 202 Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Limiting Reagent Worksheet 202 Answer Key Compatibility with Devices
 - Limiting Reagent Worksheet 202 Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limiting Reagent Worksheet 202 Answer Key
 - Highlighting and Note-Taking Limiting Reagent Worksheet 202 Answer Key
 - Interactive Elements Limiting Reagent Worksheet 202 Answer Key

8. Staying Engaged with Limiting Reagent Worksheet 202 Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limiting Reagent Worksheet 202 Answer Key
9. Balancing eBooks and Physical Books Limiting Reagent Worksheet 202 Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting Reagent Worksheet 202 Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Limiting Reagent Worksheet 202 Answer Key
 - Setting Reading Goals Limiting Reagent Worksheet 202 Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Limiting Reagent Worksheet 202 Answer Key
 - Fact-Checking eBook Content of Limiting Reagent Worksheet 202 Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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