

**STAAR CHEMISTRY
REFERENCE MATERIALS**

ATOMIC STRUCTURE

$$\text{Speed of light} = (\text{frequency})(\text{wavelength}) \quad c = f\lambda$$

$$\text{Energy} = (\text{Planck's constant})(\text{frequency}) \quad E_{\text{photon}} = hf$$

$$\text{Energy} = \frac{(\text{Planck's constant})(\text{speed of light})}{(\text{wavelength})} \quad E_{\text{photon}} = \frac{hc}{\lambda}$$

BEHAVIOR OF GASES

$$\text{Total pressure of a gas} = \left(\text{sum of the partial pressures of the component gases} \right) \quad P_T = P_1 + P_2 + P_3 + \dots$$

$$(\text{Pressure})(\text{volume}) = (\text{moles})(\text{ideal gas constant})(\text{temperature}) \quad PV = nRT$$

$$\frac{(\text{initial pressure})(\text{initial volume})}{(\text{initial moles})(\text{initial temperature})} = \frac{(\text{final pressure})(\text{final volume})}{(\text{final moles})(\text{final temperature})} \quad \frac{P_1V_1}{n_1T_1} = \frac{P_2V_2}{n_2T_2}$$

$$(\text{initial pressure})(\text{initial volume}) = (\text{final pressure})(\text{final volume}) \quad P_1V_1 = P_2V_2$$

$$\frac{(\text{initial volume})}{(\text{initial temperature})} = \frac{(\text{final volume})}{(\text{final temperature})} \quad \frac{V_1}{T_1} = \frac{V_2}{T_2}$$

$$\frac{(\text{initial volume})}{(\text{initial moles})} = \frac{(\text{final volume})}{(\text{final moles})} \quad \frac{V_1}{n_1} = \frac{V_2}{n_2}$$

SOLUTIONS

$$\text{Molarity} = \frac{\text{moles of solute}}{\text{volume of solution}} \quad M = \frac{\text{mol}}{L}$$

$$\text{Ionization constant of water} = \left(\text{hydrogen ion concentration} \right) \left(\text{hydroxide ion concentration} \right) \quad K_w = [H^+][OH^-]$$

$$\left(\text{volume of solution 1} \right) \left(\text{molarity of solution 1} \right) = \left(\text{volume of solution 2} \right) \left(\text{molarity of solution 2} \right) \quad V_1M_1 = V_2M_2$$

$$\text{pH} = -\log(\text{hydrogen ion concentration}) \quad \text{pH} = -\log[H^+]$$

THERMOCHEMISTRY

$$\text{Heat gained or lost} = (\text{mass}) \left(\text{specific heat} \right) \left(\text{change in temperature} \right) \quad Q = mc_p\Delta T$$

$$\text{Enthalpy of reaction} = \left(\text{enthalpy of products} \right) - \left(\text{enthalpy of reactants} \right) \quad \Delta H = \Delta H^\circ(\text{products}) - \Delta H^\circ(\text{reactants})$$

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