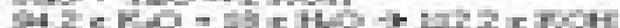


KEY

Limiting Reagents: Product Formation

1. If 20.0 grams of K_2O react with 20.0 g of H_2O , how many grams of KOH are formed?



$$K_2O \text{ Scale Factor} = \frac{20.0 \text{ g } K_2O}{94.2 \text{ g } K_2O} = 0.212$$

$$H_2O \text{ Scale Factor} = \frac{10.0 \text{ g } H_2O}{18.0 \text{ g } H_2O} = 0.556$$

$$KOH \text{ Limiting} \\ (0.212)(112.2 \text{ g KOH}) = 23.8 \text{ g KOH}$$

2. If 100.0 grams of mercury (II) hydroxide react with 50.0 grams of phosphoric acid, how many grams of mercury (II) phosphate are formed?



$$H_3PO_4 \text{ Scale Factor} = \frac{50.0 \text{ g } H_3PO_4}{196.0 \text{ g } H_3PO_4} = 0.2551$$

$$Hg_2(OH)_2 \text{ Limiting} \\ (0.1421)(794.8 \text{ g } Hg_2(PO_4)_3) = 112.3 \text{ g } Hg_2(PO_4)_3$$

3. If 20.0 grams of KOH react with 98.0 grams of phosphoric acid, how many grams of potassium phosphate are formed?



$$KOH \text{ Scale Factor} = \frac{20.0 \text{ g KOH}}{168.0 \text{ g KOH}} = 0.1194$$

$$H_3PO_4 \text{ Scale Factor} = \frac{98.0 \text{ g } H_3PO_4}{98.0 \text{ g } H_3PO_4} = 1.0000$$

$$KOH \text{ Limiting} \\ (0.1194)(212.0 \text{ g } K_3PO_4) = 25.3 \text{ g } K_3PO_4$$

4. If 15.0 grams of calcium chloride react with 15.0 grams of nitric acid, how many grams of calcium nitrate are formed? How many grams of hydrochloric acid are formed?



$$CaCl_2 \text{ Scale Factor} = \frac{15.0 \text{ g } CaCl_2}{111.0 \text{ g } CaCl_2} = 0.1351$$

$$HNO_3 \text{ Scale Factor} = \frac{15.0 \text{ g } HNO_3}{126.0 \text{ g } HNO_3} = 0.119$$

$$HNO_3 \text{ Limiting} \\ (0.119)(164.1 \text{ g } Ca(NO_3)_2) = 19.5 \text{ g } Ca(NO_3)_2 \quad (0.119)(72.9 \text{ g } HCl) = 8.68 \text{ g } HCl$$

5. If 30.0 grams of hydrogen chloride react with 40.0 grams of calcium hydroxide, how many grams of calcium chloride can be made?



Limiting And Excess Reactants Pogil Key Answers

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Limiting Reactant and Percent Yield, 2019 This lesson plan covers analyzing chemical reactions in order to determine limiting reactants and excess reactants including calculating the amount of excess reactant calculate the theoretical yield of a reaction and calculating the percent yield of a reaction

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