

## Limiting Reagent and Percentage Yield

- 1 Dichlorine monoxide is produced by passing chlorine gas over heated mercury (II) oxide according to the following unbalanced reaction:



Balance the reaction. Suppose the quantity of the reactants is sufficient to produce 0.86 g of  $\text{Cl}_2\text{O}$ , but only 0.71 g is obtained. What is the percentage yield?

- 2 Consider the following reaction:



The theoretical yield of beryllium chloride was 10.7 g. Suppose the reaction actually yields 4.5 g. What is the percentage yield?

- 3 Nitrogen and hydrogen at high temperatures are converted to ammonia using the following reaction:



When 400 g of  $\text{H}_2$  are added to an excess amount of  $\text{N}_2$ , 104 g of  $\text{NH}_3$  are formed. Calculate the percentage yield.

# Limiting Reactant And Percent Yield Review Sheet 3 Answers

**Robert R. Redfield**



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