

1. What volume of hydrogen gas at STP is produced from the reaction of 50.0g of Mg and 75.0 grams of HCl? How much of the excess reagent is left over (in grams)?



Do a standard mass to mass problem starting with each reactant

50.0 g Mg	1 mol Mg 24.31g Mg	1 mol H <sub>2</sub> 1 mol Mg	22.4 L 1 mol H <sub>2</sub>	= 46.1 L H <sub>2</sub>
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75.0 g HCl	1 mol HCl 36.46 g HCl	1 mole H <sub>2</sub> 2 mol HCl	22.4 L 1 mol H <sub>2</sub>	= 23.0 L H <sub>2</sub>
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**HCl is the limiting reactant!!**

# Limiting Reactants Pogil Key

**Puja Mehta**



**Limiting Reactants Pogil Key:**

Limiting Reactant and Percent Yield, 2019 This lesson plan covers analyzing chemical reactions in order to determine limiting reactants and excess reactants including calculating the amount of excess reactant calculate the theoretical yield of a reaction and calculating the percent yield of a reaction

## The Enigmatic Realm of **Limiting Reactants Pogil Key**: Unleashing the Language is Inner Magic

In a fast-paced digital era where connections and knowledge intertwine, the enigmatic realm of language reveals its inherent magic. Its capacity to stir emotions, ignite contemplation, and catalyze profound transformations is nothing lacking extraordinary. Within the captivating pages of **Limiting Reactants Pogil Key** a literary masterpiece penned by a renowned author, readers embark on a transformative journey, unlocking the secrets and untapped potential embedded within each word. In this evaluation, we shall explore the book's core themes, assess its distinct writing style, and delve into its lasting impact on the hearts and minds of people who partake in its reading experience.

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