

For the following reactions, find the following:

- a) Which of the reagents is the limiting reagent?
- b) What is the maximum amount of each product that can be formed?
- c) How much of the other reagent is left over after the reaction is complete?

1) Consider the following reaction:



Answer the questions above, assuming we started with 30 grams of ammonium nitrate and 50 grams of sodium phosphate.

30.0 g NH <sub>4</sub> NO <sub>3</sub>	1 mole	3 moles NaNO <sub>3</sub>	85.00 g NaNO <sub>3</sub>	-31.9 g NaNO <sub>3</sub> <i>max amount!</i>
80.06 g	3 moles NH <sub>4</sub> NO <sub>3</sub>	1 mole NaNO <sub>3</sub>	85.00 g NaNO <sub>3</sub>	

limiting reactant is NH<sub>4</sub>NO<sub>3</sub>

50 g Na <sub>3</sub> PO <sub>4</sub>	1 mole Na <sub>3</sub> PO <sub>4</sub>	3 moles NaNO <sub>3</sub>	85.00 g NaNO <sub>3</sub>	= 77.8 g NaNO <sub>3</sub>
163.94 g	1 mole Na <sub>3</sub> PO <sub>4</sub>	1 mole NaNO <sub>3</sub>	1 mole NaNO <sub>3</sub>	

ⓐ 30.0 g NH<sub>4</sub>NO<sub>3</sub>

1 mole NH <sub>4</sub> NO <sub>3</sub>	1 mole Na <sub>3</sub> PO <sub>4</sub>	163.94 g	20.5 g Na <sub>3</sub> PO <sub>4</sub> needed
80.06 g NH <sub>4</sub> NO <sub>3</sub>	3 moles NH <sub>4</sub> NO <sub>3</sub>	1 mole Na <sub>3</sub> PO <sub>4</sub>	

left over: 50.0 - 20.5 = 29.5 g left over  
start used

# Limiting Reagent Test 2014 Form Answers

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1. Understanding the eBook Limiting Reagent Test 2014 Form Answers
  - The Rise of Digital Reading Limiting Reagent Test 2014 Form Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reagent Test 2014 Form Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Limiting Reagent Test 2014 Form Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reagent Test 2014 Form Answers

- Personalized Recommendations
  - Limiting Reagent Test 2014 Form Answers User Reviews and Ratings
  - Limiting Reagent Test 2014 Form Answers and Bestseller Lists
5. Accessing Limiting Reagent Test 2014 Form Answers Free and Paid eBooks
- Limiting Reagent Test 2014 Form Answers Public Domain eBooks
  - Limiting Reagent Test 2014 Form Answers eBook Subscription Services
  - Limiting Reagent Test 2014 Form Answers Budget-Friendly Options
6. Navigating Limiting Reagent Test 2014 Form Answers eBook Formats
- ePub, PDF, MOBI, and More
  - Limiting Reagent Test 2014 Form Answers Compatibility with Devices
  - Limiting Reagent Test 2014 Form Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Limiting Reagent Test 2014 Form Answers
  - Highlighting and Note-Taking Limiting Reagent Test 2014 Form Answers
  - Interactive Elements Limiting Reagent Test 2014 Form Answers
8. Staying Engaged with Limiting Reagent Test 2014 Form Answers
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Limiting Reagent Test 2014 Form Answers
9. Balancing eBooks and Physical Books Limiting Reagent Test 2014 Form Answers
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Limiting Reagent Test 2014 Form Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Limiting Reagent Test 2014 Form Answers
- Setting Reading Goals Limiting Reagent Test 2014 Form Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Limiting Reagent Test 2014 Form Answers

- 
- Fact-Checking eBook Content of Limiting Reagent Test 2014 Form Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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