

CHEMICAL EQUILIBRIUM WORKSHEET

On the line at the left, write the letter of the description that best matches each term.

- | | | |
|-------|--------------------------------|--|
| _____ | 1. Equilibrium position | a. used to determine if a reaction has reached equilibrium |
| _____ | 2. Law of chemical equilibrium | b. depends on the initial concentrations of the substances in a reaction |
| _____ | 3. Reaction quotient | c. states that every reaction proceeds to an equilibrium state with a specific K_{eq} |
| _____ | 4. Law of mass action | d. expresses the relative concentration of reactants and products at equilibrium in terms of an equilibrium constant |
| _____ | 5. Equilibrium constant | e. the ratio of product concentration to reactant concentration at equilibrium |

Answer each of the following in the space provided

6. What is the equilibrium expression for the equation $H_2(g) + I_2(g) \rightleftharpoons 2HI(g)$?
7. What is the equilibrium expression for the equation $NH_4Cl(s) \rightleftharpoons NH_3(g) + HCl(g)$?
8. What is the equilibrium expression for the equation $As_4O_6(s) + 6C(s) \rightleftharpoons As_4(g) + 6CO(g)$?
9. What is the equilibrium expression for the equation $SnO_2(s) + 2CO(g) \rightleftharpoons Sn(s) + 2CO_2(g)$?
10. What is the equilibrium expression for the equation $CaCO_3(s) \rightleftharpoons CaO(s) + CO_2(g)$?
11. For the reaction $2CO(g) \rightleftharpoons C(s) + CO_2(g)$, $K_{eq} = 7.7 \times 10^{-28}$. At a particular time, the following concentrations are measured: $[CO]=0.034$ M, $[CO_2]=3.6 \times 10^{-17}$ M. Is this reaction at equilibrium? If not which direction will the reaction proceed?
12. For the reaction $N_2O_4(g) \rightleftharpoons 2NO_2(g)$, $K_{eq} = 0.2$. At a particular time, the following concentrations are measured: $[N_2O_4]=2.0$ M, $[NO_2] = 0.2$ M. Is this reaction at equilibrium? If not which direction will the reaction proceed?
13. For the reaction $2HCl(g) \rightleftharpoons H_2(g) + Cl_2(g)$, $K_{eq} = 0.11$. At a particular time, the following concentrations are measured: $[HCl]=2.5$ M, $[H_2] = 2.0$ M, $[Cl_2]= 1.2$ M. Is this reaction at equilibrium? If not which direction will the reaction proceed?
14. At 340 °C, $K_{eq} = 0.064$ for the reaction $Fe_2O_3(s) + 3H_2(g) \rightleftharpoons 2Fe(s) + 3H_2O(g)$ Given that $[H_2]=0.45$ M and $[H_2O]=0.37$ M, find Q and predict how the reaction will proceed.

Match each statement with the appropriate letter. Each letter can be used once, more than once, or not at all.

- | | | |
|-------|---|-------------------------------------|
| _____ | 15. The equilibrium concentration of products is much greater than that of reactants. | a. K_{eq} is much greater than 1. |
| _____ | 16. The equilibrium concentration of products is much less than that of reactants | b. K_{eq} is about equal to 1. |
| _____ | 17. There is a considerable amount of both reactants and products at equilibrium | c. K_{eq} is much less than 1. |

Holt Chemistry Chemical Equilibrium Answer Key Test

Michael Hoelscher



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