



Covalent Compounds Unit Study Guide

Name _____

Teacher _____

Date: _____

Period _____

Reference:

Chemistry in the Community, pp 181-183,
Notes
Naming covalent compounds worksheet
Viscosity Lab
Exploring Properties of Ionic and Covalent compounds

Vocabulary

Understand and apply the definitions of the following terms:

Covalent

Valence electrons

e dot diagrams (Lewis dot diagrams)

Polar

Nonpolar

Intermolecular forces

Octet rule

Polymers

Viscosity

Boiling points

Targeted Skills

- Diagram a covalent molecule using a e dot diagrams (Lewis dot diagrams).
- Know the following properties:
 - > Ionic compounds = the transfer of electron(s), i.e. "gain or lose"
Metal cations & nonmetal anion (or polyatomic ion)
Net charge = zero (criss cross charges) $X^{+n}Y^{-m} = X_nY_m$
Dissociate (dissolve) in water
 - > Molecular compounds = sharing electrons by forming a covalent bond
2 or more nonmetals
Prefixes indicate # of each atom
Polar molecules have a δ^+ & a δ^-
Polar molecules dissolve in water
Nonpolar molecules do not have oppositely charged ends, therefore do not dissolve in water.
 - > Octet rule = atoms share, gain, or lose e- to obtain full outer energy level (usually 8 e-)
- Name covalent compounds using prefixes: (mono, di, tri, tetra, penta, hexa, hepta, octa, nona, deca).
- Compare/contrast the arrangement of atoms in molecules to the arrangement of atoms in ionic compounds.
- Describe the influence of intermolecular forces on physical and chemical properties of covalent compounds. Example: Boiling points, viscosity, and density.
- Be able to describe properties (boiling points, melting points, viscosity, solubility, conductivity) of polymers, ionic crystals, metallic substances, atoms and molecules.

Make sure you remember from past tests?

Atomic structure

Significant figures

Accuracy & precision

Mixtures

Nomenclature

Solubility

Safety

Elements

Ionic compound formula writing

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Ralph Thomas Myers



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