

Intro to Energy, Enzymes, and Metabolism

Sept 24th

Concepts in Energy and Thermodynamics

- Energy
- Reactions
- Entropy
- 2nd law of thermodynamics
- Free energy
- Kinetics of chemical reactions
- ATP
- Catalysis and Enzymes

Energy is the ability to do work, and it will determine the **direction** and **rate** of reaction ($aA + bB \leftrightarrow cC + dD$)

Enthalpy (H)

- The total energy of a system

Free Energy (G)

- The amount of a systems energy that is available and can be used to promote change or do

Entropy (S)

- The degree of disorder of a system
- Higher Entropy → More Disordered
- Lower Entropy → Highly Ordered

T → Temperature

Total Energy = Usable Energy + Unusable Energy

$$(H = G + TS)$$

$$(\Delta G = \Delta H - T \Delta S)$$

- Negative Delta G → loss of energy = Spontaneous reaction

2nd Law of Thermodynamics

- The state of a natural system itself can be reserved, but **not** without increasing the entropy of the systems surroundings
- The change in free energy determines reaction direction
- Spontaneous doesn't need a input of energy
- Delta G will tell you what direction it will go, not the rate of the reaction

Mastering Biology Ch6

Lingsheng Yao



Mastering Biology Ch6:

Reviewing **Mastering Biology Ch6**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**Mastering Biology Ch6**," an enthralling opus penned by a very acclaimed wordsmith, readers set about an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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