

Part A

Calculate the amount of heat (in kilojoules) required to vaporize 2.58 kg of water at its boiling point. Express the heat in kilojoules to three significant figures.

$$q = 5830 \text{ kJ}$$

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Correct

At its boiling point, the enthalpy of vaporization (ΔH_{vap}) of water is 40.7 kJ/mol. Once the given mass, 2.58×10^3 g, is converted to the number of moles using the molar mass of water (18.02 g/mol), multiplying the number of moles by ΔH_{vap} yields the amount of heat required to evaporate that amount of water.

Part B

Suppose that 0.48 g of water at 25°C condenses on the surface of a 55-g block of aluminum that is initially at 25°C. If the heat released during condensation goes only toward heating the metal, what is the final temperature (in degrees Celsius) of the metal block? (The specific heat capacity of aluminum, $C_{s,\text{Al}}$, is 0.903 J/(g · °C).)

Express the temperature in degrees Celsius to two significant figures.

$T_f =$

°C

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John Olmsted, Gregory M. Williams



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Table of Contents Mastering Chemistry Answers Ch 12

1. Understanding the eBook Mastering Chemistry Answers Ch 12
 - The Rise of Digital Reading Mastering Chemistry Answers Ch 12
 - Advantages of eBooks Over Traditional Books
2. Identifying Mastering Chemistry Answers Ch 12
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mastering Chemistry Answers Ch 12
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mastering Chemistry Answers Ch 12
 - Personalized Recommendations
 - Mastering Chemistry Answers Ch 12 User Reviews and Ratings

- Mastering Chemistry Answers Ch 12 and Bestseller Lists
- 5. Accessing Mastering Chemistry Answers Ch 12 Free and Paid eBooks
 - Mastering Chemistry Answers Ch 12 Public Domain eBooks
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 - Mastering Chemistry Answers Ch 12 Budget-Friendly Options
- 6. Navigating Mastering Chemistry Answers Ch 12 eBook Formats
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 - Mastering Chemistry Answers Ch 12 Compatibility with Devices
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- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mastering Chemistry Answers Ch 12
 - Highlighting and Note-Taking Mastering Chemistry Answers Ch 12
 - Interactive Elements Mastering Chemistry Answers Ch 12
- 8. Staying Engaged with Mastering Chemistry Answers Ch 12
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mastering Chemistry Answers Ch 12
- 9. Balancing eBooks and Physical Books Mastering Chemistry Answers Ch 12
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mastering Chemistry Answers Ch 12
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mastering Chemistry Answers Ch 12
 - Setting Reading Goals Mastering Chemistry Answers Ch 12
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mastering Chemistry Answers Ch 12
 - Fact-Checking eBook Content of Mastering Chemistry Answers Ch 12
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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