

**Review Test 1**

1. What is the number of moles in a 90.20-gram sample of  $\text{MgCO}_3$  (molecular weight = 84.3 g/mol)?

- a. 0.949 mol
- b. 0.935 mol
- c. 0.843 mol
- d. 1.070 mol
- e. 1.230 mol

2. Which phrase describes the nucleus of any atom?

- a. contains negative electrons
- b. contains positive electrons
- c. has an overall positive charge
- d. has an overall neutral charge
- e. has an overall negative charge

3. Which two particles each have a mass of approximately one atomic mass unit?

- a. a neutron and an electron
- b. a proton and a neutron
- c. an electron and a proton
- d. a neutron and a positron
- e. an electron and a positron

4. Which statement describes a chemical property of iron?

- a. Iron is a solid at room temperature
- b. Iron conducts electricity.
- c. Iron is malleable.
- d. Iron reacts with nitric acid
- e. Iron has a high melting point.

5. Which substance can be broken down by chemical change?

- a. cyclohexane
- b. copper
- c. aluminum
- d. magnesium
- e. neon

6. Which process can be used to separate a mixture of two liquids having different boiling points?

- a. crystallization
- b. deposition
- c. filtration
- d. distillation
- e. sublimation

7. Which statement describes a solution of sodium chloride in water?

- a. The mixture is a colloid, the solute is  $\text{H}_2\text{O}$  and the solvent is  $\text{NaCl}$ .
- b. The mixture is heterogeneous, the solute is  $\text{H}_2\text{O}$  and the solvent is  $\text{NaCl}$ .
- c. The mixture is homogeneous, the solute is  $\text{NaCl}$  and the solvent is  $\text{H}_2\text{O}$ .
- d. The mixture is homogeneous, the solute is  $\text{H}_2\text{O}$  and the solvent is  $\text{NaCl}$ .
- e. The mixture is heterogeneous, the solute is  $\text{NaCl}$  and the solvent is  $\text{H}_2\text{O}$ .

8. A potassium atom has a mass number of 41. What is the number of neutrons in this atom?

- a. 15
- b. 22
- c. 20
- d. 41
- e. 21

9. How many electrons are contained in a  $\text{K}^{+1}$  ion?

- a. 17
- b. 18
- c. 19
- d. 20
- e. 21

# Gradpoint Chemistry Review Test

**S Ashworth**

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