

# Graphing Periodic Trends Lab

Name \_\_\_\_\_ Hr \_\_\_\_\_

**PURPOSE:** To determine how certain properties are periodic when the elements are arranged in periods or groups.

**DISCUSSION:** There are many periodic trends that occur as you cross a period or go down a group/family on the Periodic Table. Graphing various elemental properties will help to illustrate these trends.

## REFERENCE:

Angstrom - A unit of distance equal to one ten-billionth of a meter.

Atomic Radius - The distance from the center of the nucleus to the valence electron energy level.

Ionization energy - The energy required to remove an electron from an atom.

## DATA

**Properties of the First 18 Elements**

| Element | Atomic Number | Atomic Radius (Angstroms) | Ionization Energy |
|---------|---------------|---------------------------|-------------------|
| H       | 1             | 0.79                      | 314               |
| He      | 2             | 0.49                      | 567               |
| Li      | 3             | 2.05                      | 124               |
| Be      | 4             | 1.4                       | 215               |
| B       | 5             | 1.17                      | 191               |
| C       | 6             | 0.91                      | 260               |
| N       | 7             | 0.75                      | 335               |
| O       | 8             | 0.65                      | 314               |
| F       | 9             | 0.57                      | 402               |
| Ne      | 10            | ?                         | ?                 |
| Na      | 11            | 2.23                      | 119               |
| Mg      | 12            | 1.72                      | 176               |
| Al      | 13            | 1.82                      | 138               |
| Si      | 14            | 1.46                      | 168               |
| P       | 15            | 1.23                      | 242               |
| S       | 16            | 1.09                      | 239               |
| Cl      | 17            | 0.97                      | 299               |
| Ar      | 18            | ?                         | ?                 |

Using the data above, make a line graph the following data. Draw a dark vertical line to separate the 3 periods. Give each graph a title and label both the x-axis and the y-axis. Connect your points. Predict the plot points for elements with question marks.

1. Atomic Radius (y) vs. Atomic Number (x) for the first 3 periods (18 elements – 3 graphs)
2. Ionization energy (y) vs. Atomic Number (x) for the first 3 periods (18 elements – 3 graphs)

# Graphing Trends In The Periodic Table Lab Answers Key

**JS Bruner**



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