

# JEE Main 2013 Question Paper with Solutions

## Test Booklet Code Q

### PART A – CHEMISTRY

1. An unknown alcohol is treated with the "Lucas reagent" to determine whether the alcohol is primary, secondary or tertiary. Which alcohol reacts fastest and by what mechanism?
- (1) tertiary alcohol by  $S_N1$  (2) secondary alcohol by  $S_N2$   
(3) tertiary alcohol by  $S_N2$  (4) secondary alcohol by  $S_N1$

**Sol.** (1)  
Reaction proceeds through carbocation formation as  $3^\circ$  carbocation is highly stable, hence reaction proceeds through  $S_N1$  with  $3^\circ$  alcohol.

2. The first ionization potential of Na is 5.1 eV. The value of electron gain enthalpy of  $Na^+$  will be:
- (1)  $-5.1$  eV (3)  $-10.2$  eV  
(2)  $+2.55$  eV (4)  $-2.55$  eV

**Sol.** (1)  
 $Na \xrightarrow[\Delta H = -5.1 eV]{\Delta H = 5.1 eV} Na^+ + e^-$ , here the backward reaction releases same amount of energy and known as Electron gain enthalpy.

3. Stability of the species  $Li_2$ ,  $Li_2^+$  and  $Li_2^-$  increases in the order of:
- (1)  $Li_2^+ < Li_2^- < Li_2$  (3)  $Li_2 < Li_2^+ < Li_2^-$   
(2)  $Li_2^- < Li_2 < Li_2^+$  (4)  $Li_2 < Li_2^- < Li_2^+$

**Sol.** (1)  
 $Li_2 (6) = \sigma 1s^2 \ \delta 1s^2 \ \sigma 2s^2$   
B.O. =  $\frac{4-2}{2} = 1$   
 $Li_2^+ (5) = \sigma 1s^2 \ \delta 1s^1 \ \sigma 2s^1$   
B.O. =  $\frac{3-2}{2} = 0.5$   
 $Li_2^- (7) = \sigma 1s^2 \ \delta 1s^2 \ \sigma 2s^2 \ \delta 2s^1$   
B.O. =  $\frac{4-3}{2} = 0.5$   
 $Li_2^+$  is more stable than  $Li_2^-$  because  $Li_2^+$  has more numbers of antibonding electrons.

4. The molarity of a solution obtained by mixing 750 mL of 0.5 (M) HCl with 250 mL of 2(M)HCl will be:
- (1) 1.00 M (2) 1.75 M  
(3) 0.975 M (4) 0.875 M

**Sol.** (4)  
 $M_1V_1 + M_2V_2 = MV$   
 $M = \frac{M_1V_1 + M_2V_2}{V} = \frac{0.5 \times 750 + 2 \times 250}{1000}$   
 $M = 0.875$

5. Which of the following is the wrong statement?
- (1)  $O_3$  molecule is bent (2) Ozone is violet-black in solid state  
(3) Ozone is diamagnetic gas (4)  $ONCl$  and  $ONF$  are not isoelectronic

**Sol.** (All the options are correct statements)

(1) Correct, as  is bent.

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**Shasha Hu**



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