

# General Chemistry Homework Answer Key

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Textbook: OpenStax Chemistry: Atoms First 2e

Chapter 3: Electronic Structure and Periodic Properties of Elements

TA Reference Answers with Step-by-Step Explanations

## 3.1 Electromagnetic Energy

- 2 – Define frequency, wavelength, and amplitude.

Answer:

Frequency ( $\nu$ ) is the number of wave cycles that pass a point per second, measured in hertz (Hz).

Wavelength ( $\lambda$ ) is the distance between successive crests of a wave, measured in meters (m).

Amplitude is the height of the wave from the center line to the crest or trough, indicating intensity.

Stepwise Procedure:

1. Define each term conceptually.
2. Include appropriate units.
3. Clarify differences among the properties.

- 4 – What is the relationship between frequency and energy?

Answer:

The energy ( $E$ ) of electromagnetic radiation is directly proportional to its frequency ( $\nu$ ):  $E = h\nu$ , where  $h$  is Planck's constant ( $6.626 \times 10^{-34}$  J·s).

Stepwise Procedure:

1. State the energy-frequency equation.
2. Define  $h$  and  $\nu$ .
3. Explain that higher frequency means higher energy.

## 3.2 The Bohr Model

- 8 – What are the key features of the Bohr model of the hydrogen atom?

Answer:

1. Electrons orbit the nucleus in fixed paths (energy levels).
2. Only certain orbits are allowed with quantized energies.
3. Electrons absorb or emit energy when they move between levels.

# General Chemistry Homework Guide

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