

NAME \_\_\_\_\_

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### GRAPHING PERIODIC TRENDS

The Periodic Table is arranged according to the Periodic Law. The Periodic Law states that when elements are arranged in order of increasing atomic number, their physical and chemical properties show a periodic pattern.

Define the following terms(starting on page 170):

Atomic radius-

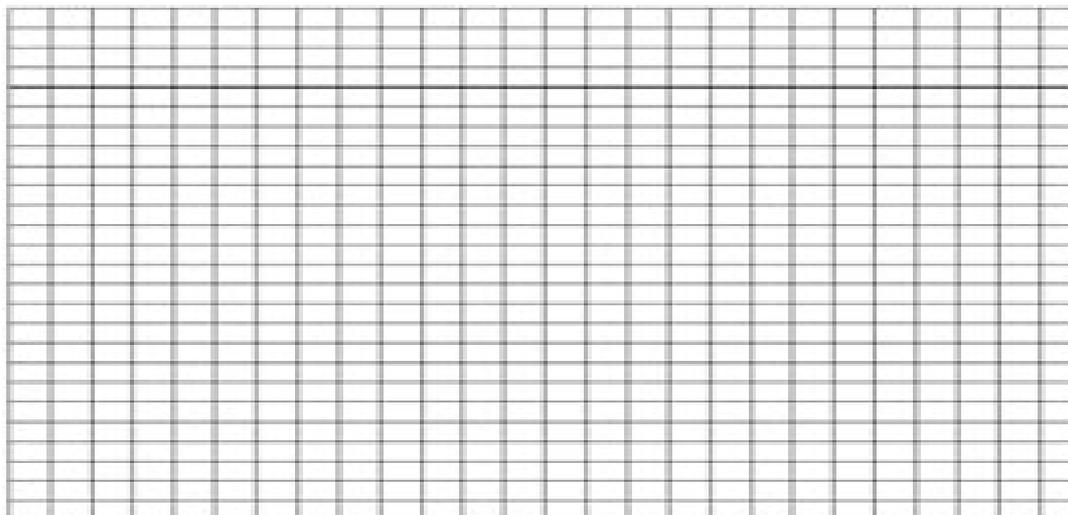
First ionization energy-

Electronegativity-

#### Graph 1 – Atomic Radius as a function of Atomic Number

- A. Create a graph of the atomic radius as a function of atomic number. Plot atomic number on the X axis and atomic radius on the Y axis. Remember to label the axes!
- B. Use a colored pen, pencil or highlighter to trace over the element's period (horizontal row on the periodic table). For example: use GREEN to trace for all of the elements in row 1, then use YELLOW to trace for all of the elements in row 2, then use ORANGE to trace for all the elements in row 3, then use BLUE to trace for all the elements in row 4.

Symbol	Atomic Number	Atomic Radius(pm)
H	1	37
He	2	50
Li	3	123
Be	4	89
B	5	80
C	6	77
N	7	70
O	8	66
F	9	64
Ne	10	67
Na	11	157
Mg	12	136
Al	13	125
Si	14	117
P	15	110
S	16	104
Cl	17	99
Ar	18	98
K	19	203
Ca	20	174



- Describe the trend in atomic radius as the atomic number increases across a period.
- Describe the trend in atomic radius as the atomic number increases down a group.

# Graphing Periodic Trends Activity

**Michael Hoelscher**



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## **Table of Contents Graphing Periodic Trends Activity**

1. Understanding the eBook Graphing Periodic Trends Activity
  - The Rise of Digital Reading Graphing Periodic Trends Activity
  - Advantages of eBooks Over Traditional Books
2. Identifying Graphing Periodic Trends Activity
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Graphing Periodic Trends Activity
  - User-Friendly Interface
4. Exploring eBook Recommendations from Graphing Periodic Trends Activity
  - Personalized Recommendations
  - Graphing Periodic Trends Activity User Reviews and Ratings
  - Graphing Periodic Trends Activity and Bestseller Lists
5. Accessing Graphing Periodic Trends Activity Free and Paid eBooks
  - Graphing Periodic Trends Activity Public Domain eBooks
  - Graphing Periodic Trends Activity eBook Subscription Services

- Graphing Periodic Trends Activity Budget-Friendly Options
- 6. Navigating Graphing Periodic Trends Activity eBook Formats
  - ePub, PDF, MOBI, and More
  - Graphing Periodic Trends Activity Compatibility with Devices
  - Graphing Periodic Trends Activity Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Graphing Periodic Trends Activity
  - Highlighting and Note-Taking Graphing Periodic Trends Activity
  - Interactive Elements Graphing Periodic Trends Activity
- 8. Staying Engaged with Graphing Periodic Trends Activity
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Graphing Periodic Trends Activity
- 9. Balancing eBooks and Physical Books Graphing Periodic Trends Activity
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Graphing Periodic Trends Activity
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Graphing Periodic Trends Activity
  - Setting Reading Goals Graphing Periodic Trends Activity
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Graphing Periodic Trends Activity
  - Fact-Checking eBook Content of Graphing Periodic Trends Activity
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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