

Reading: The Particle Theory of Matter

Matter makes up everything in the universe. It is defined as anything that has **mass** and occupies **space**. The air we inhale, the chair we sit on, and the water we drink are all examples of matter.

Scientists rely on the **Particle Theory of Matter** to describe and predict how matter behaves. According to this theory, all matter is composed of **extremely small particles** that are in continuous motion. The distance between these particles and the strength of the forces holding them together determine whether a substance is a solid, liquid, or gas. When energy in the form of heat is added, particles move faster, often causing changes in state such as melting, evaporation, or boiling.



Solid



Liquid



Gas

The Particle Theory explains the differences among the **states of matter**. In a **solid**, particles are tightly packed and vibrate in fixed positions, giving solids a definite shape and volume. In a **liquid**, particles remain close together but are able to slide past one another. Liquids therefore maintain a constant volume but adopt the shape of their container. In a **gas**, particles are far apart and move independently. Gases expand to fill both the shape and the volume of their container.

Everyday experiences provide clear evidence of this theory. Ice cubes melt when energy from heat increases the motion of water particles. Sugar dissolves faster in hot tea than in iced tea because higher temperatures increase particle motion and collisions. The scent of perfume or popcorn spreads across a room because gas particles are constantly in motion, diffusing throughout the air.

The 5 Postulates of the Particle Theory of Matter

- All matter is composed of tiny particles.
- All particles are in constant motion.
- Spaces exist between particles.
- Particles exert attractive forces on one another.
- Adding heat increases particle motion.

Introduction To Matter Guided Reading And Study Key

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