

Entropy WS 11

Section A

- 1 The equation for the reaction of CO_2 and H_2O to produce glucose, $\text{C}_6\text{H}_{12}\text{O}_6$, and O_2 is shown below.



Standard entropies are given in the table below.

Substance	$\text{CO}_2(\text{g})$	$\text{H}_2\text{O}(\text{l})$	$\text{O}_2(\text{g})$
$S^\circ / \text{J K}^{-1} \text{ mol}^{-1}$	214	70	205

- (a) (i) Calculate the standard entropy of glucose.

$$\begin{aligned} 256 &= [n + 6(205)] - [6(214) + 6(70)] \\ -256 &= (n + 1230) - (1284 + 420) \\ -256 &= n + 1230 - 1704 \\ n &= -256 + 474 \\ &= +218 \text{ J K}^{-1} \text{ mol}^{-1} \end{aligned}$$

$$S^\circ = \dots\dots\dots +218 \dots\dots\dots \text{ J K}^{-1} \text{ mol}^{-1} \quad [2]$$

- (ii) Calculate ΔG , in kJ mol^{-1} , at 25°C .

Show all your working.

$$\begin{aligned} \Delta G &= \Delta H - T\Delta S \\ &= +2879 - 298 \left(\frac{-256}{1000} \right) \\ &= 2879 + 76.28 \\ &= 2955.28 \end{aligned}$$

$$\Delta G = \dots\dots\dots +2955.28 \dots\dots\dots \text{ kJ mol}^{-1} \quad [2]$$

- (iii) Explain why this reaction is not feasible at any temperature.

As ΔH is endothermic and ΔS is negative, the value for ΔG will always be positive. Therefore this reaction will not be feasible at any temperature.

[1]

Entropy Section Review Answers

Michael Brown



Entropy Section Review Answers:

Applied Mechanics Reviews, 1986 *Handbook of Mathematical Fluid Dynamics* S. Friedlander, D. Serre, 2002-07-09

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Proceedings ,2004 **Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl** Paul B. Kelter,1993 **Modern Chemistry** Holt Rinehart & Winston,Holt, Rinehart and Winston Staff,2001

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Table of Contents Entropy Section Review Answers

1. Understanding the eBook Entropy Section Review Answers
 - The Rise of Digital Reading Entropy Section Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Entropy Section Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Entropy Section Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Entropy Section Review Answers

- Personalized Recommendations
 - Entropy Section Review Answers User Reviews and Ratings
 - Entropy Section Review Answers and Bestseller Lists
5. Accessing Entropy Section Review Answers Free and Paid eBooks
 - Entropy Section Review Answers Public Domain eBooks
 - Entropy Section Review Answers eBook Subscription Services
 - Entropy Section Review Answers Budget-Friendly Options
 6. Navigating Entropy Section Review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Entropy Section Review Answers Compatibility with Devices
 - Entropy Section Review Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Entropy Section Review Answers
 - Highlighting and Note-Taking Entropy Section Review Answers
 - Interactive Elements Entropy Section Review Answers
 8. Staying Engaged with Entropy Section Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Entropy Section Review Answers
 9. Balancing eBooks and Physical Books Entropy Section Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Entropy Section Review Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Entropy Section Review Answers
 - Setting Reading Goals Entropy Section Review Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Entropy Section Review Answers

- Fact-Checking eBook Content of Entropy Section Review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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