

## UNIT III - THERMOCHEMISTRY

1. Calculate the H-F bond energy given the following data.



$$\Delta H_{\text{rxn}} = -544 \text{ kJ}$$

- A. 1130 kJ/mol    B. 565 kJ/mol    **C. 283 J/mol**    D. 300 kJ/mol    E. 134 kJ/mol
2. 1.0 g of the rocket fuel hydrazine ( $\text{N}_2\text{H}_4$ ) is burned in a bomb calorimeter. The temperature of the bomb calorimeter changes from 24.63 °C to 28.16 °C. The heat capacity of the bomb calorimeter is 837 J/K. Calculate the heat of combustion per mol of hydrazine in the bomb calorimeter.
- A. 95.0 kJ/mol**    B. 2.96 kJ/mol    C.  $3.55 \times 10^3$   
D. 92.4 J/mol    E.  $1.13 \times 10^5$  kJ/mol
3. Calculate the enthalpy change for the following reaction of nitric oxide NO with ozone  $\text{O}_3$ .
- $$\text{NO(g)} + \text{O}_3\text{(g)} \rightarrow \text{NO}_2\text{(g)} + \text{O}_2\text{(g)}$$
- Given the following data:
- |   |                                |
|---|--------------------------------|
| $\text{NO (g)} + \text{O (g)} \rightarrow \text{NO}_2 \text{(g)}$ | $\Delta H = -304.1 \text{ kJ}$ |
| $2\text{O}_3 \text{(g)} \rightarrow 3\text{O}_2 \text{(g)}$       | $\Delta H = -284.6 \text{ kJ}$ |
| $2\text{O (g)} \rightarrow \text{O}_2 \text{(g)}$                 | $\Delta H = -495.0 \text{ kJ}$ |
- A. -1083.7 kJ    B. -693.8 kJ    C. -198.8 kJ    **D. -93.7 kJ**    E. -93.7 kJ
4. An endothermic reaction does no work on the surroundings. What values do you predict for the heat and energy change of this reaction?
- A.  $q < 0, \Delta E > 0$     **B.  $q < 0, \Delta E < 0$**     C.  $q > 0, \Delta E > 0$   
D.  $q > 0, \Delta E < 0$     E.  $q > 0, \Delta E = 0$
5. M&M candies contain 209.0 Calories per 1.5 oz serving. The fat content accounts for 38.3% of these Calories. What is the energy of the fat content in kJ?
- A. 19.1 kJ    B. 874 kJ    **C.  $1.91 \times 10^{-2}$  kJ**  
D.  $3.35 \times 10^5$  kJ    E. 335 kJ
6. Calculate the heat released when 2.00 g of butane from a cigarette lighter is burned.
- $$2 \text{C}_4\text{H}_{10} \text{(l)} + 13 \text{O}_2 \text{(g)} \rightarrow 8 \text{CO}_2 \text{(g)} + 10 \text{H}_2\text{O (l)}$$
- A. -90.7 kJ    **B. -98.3 kJ**    C. -49.2 kJ    D. -108.4 kJ    E. -196.6 kJ
7. When Karl Kaveman adds chilled grog to a granite mug, 10.9 kJ of energy is transferred from the mug to the grog. If the mug has a mass of 625 g and an initial temperature of 25 °C, what is the final temperature of the mug? The specific heat of granite is 0.79 J/g·K.
- A. 2.9 °C    B. 14 °C    **C. 22 °C**    D. 25 °C    E. 47 °C

# General Chemistry Multiple Choice Questions And Answers

**Rachel S Tattersall**



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