

THE ULTIMATE GENERAL CHEMISTRY 1 STUDY GUIDE

❖ Matter and Measurement

- ✓ Matter: anything that has mass and occupies space
- ✓ Properties of matter: physical properties and chemical properties
- ✓ States of matter: solid, liquid, gas
- ✓ Measurements: SI Units, significant figures, dimensional analysis

❖ Atoms, Molecules, and Ions

- ✓ Atomic structure: protons, neutrons, and electrons
- ✓ Periodic table and trends: electronegativity, ionization energy, electron affinity
- ✓ Chemical bonding: ionic bonding, covalent bonding, metallic bonding
- ✓ Chemical formulas and equations: empirical and molecular formulas, balancing chemical equations

❖ Stoichiometry

- ✓ Limiting reactants and percent yield
- ✓ Mole concept and Avogadro's number
- ✓ Molarity and molality
- ✓ Gas laws: Boyle's law, Charles's law, Gay-Lussac's law, Avogadro's law, Ideal gas law

❖ Chemical Reactions and Equations

- ✓ Types of chemical reactions: synthesis, decomposition, single replacement, double replacement, combustion
- ✓ Redox reactions and balancing redox equations
- ✓ Reaction rates: factors affecting reaction rates, rate laws, and reaction mechanisms

❖ Gases

- ✓ Gas laws: Boyle's law, Charles's law, Gay-Lussac's law, Avogadro's law, Ideal gas law
- ✓ Kinetic Molecular Theory of gases: pressure, temperature, volume, and the number of moles of gas particles
- ✓ Natural gases: van der Waals equation

General Chemistry Thinking Guide

**McMurry, DonnaJean Fredeen, Robert
Pribush**



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