

## Chapter 21: The Evolution of Populations

This chapter begins with the idea that we focused on as we closed the last chapter: Individuals do not evolve! Populations evolve.

**Study Tip:** Always study the figures and pictures and their legends carefully as they have been chosen to illustrate and explain important ideas.

1. What are the three main mechanisms that can cause the evolution of populations?

*Natural selection, genetic drift, and gene flow are mechanisms that can cause the evolution of population.*

2. What is microevolution?

*Microevolution is a change in allele frequencies in a population over generations.*

2. Of the three mechanisms of evolution, you cited in your response to question 1, what is the only mechanism that is *adaptive*, or improves the match between organisms and their environment?

*Natural selection is an adaptive mechanism where the organisms that are better adapted to their environment tend to survive and produce more offspring.*

3. Perhaps because we tend to focus on mutations that cause changes in phenotypes, it is easy to think that every mutation will lead to a phenotypic change. Explain how a substitution error in an exon could have no effect on the amino acid sequence.

*This is because only the genetic change will be passed on to the offspring not the phenotypic change. An example of this could be in a body builder. The muscles are not passed down from generation to generation.*

4. Several sources of genetic variation are available. What is the ultimate source of new

# Evolution Of Populations Chapter Guided Reading

**M Woodhall**



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a net ionic equation is the most accurate representation of the actual chemical process that occurs writing these equations requires a familiarity with solubility rules acid base reactivity

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction

therefore the net ionic equation is  $2\text{cl aq} + 2\text{ag aq}$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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aug 10 2023 write the full ionic and net ionic equations for this reaction solution the elemental metals magnesium on the reactant side copper on the product side are neutral

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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highlights learning objectives by the end of this section you will be able to derive chemical equations from narrative descriptions of chemical reactions write and balance chemical

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the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction notice that in writing the net ionic

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what remains when the spectator ions are removed is called the net ionic equation which represents the actual chemical change occurring between the ionic compounds  $\text{Cl}^-$  aq

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver  $\text{I}^+$  ions regardless of the source of these ions these molecular and

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4 13  $2\text{Ag}^+ \text{aq} + 2\text{F}^- \text{aq} + 2\text{NH}_4^+ \text{aq} + \text{Cr}_2\text{O}_7^{2-} \text{aq} \rightarrow \text{Ag}_2\text{Cr}_2\text{O}_7 \text{s} + 2\text{NH}_4\text{F} \text{aq}$  because two  $\text{NH}_4^+$  aq

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