EXERCISE 1

[NCERT TEXT BOOK EXERCISE]

1. What are the dimensions of Planck constant? What other physical quantity has the same dimension.

2. What experimental support is available for deBroglie's concept?

3. Two particles A and B are in motion. If the wavelength associated with particle A is 5×10^{-8} m, Calculate the wavelength associated with particle B if its momentum is half of A.

4. Calculate the wavelength of 1000 kg rocket moving with a velocity of 3000 km per hour ($h = 6.626 \times 10^{-34}$ Js)

5. The sodium flame test has a characteristic yellow colour due to emissions of wavelengths 589 nm. What is the mass equivalence of one photon of this wavelength?

6. Calculate the uncertainty in the velocity of a wagon of mass 2000 kg whose position is known to an accuracy of ± 10m.

7. Calculate the uncertainty in position of a dust particle with mass equal to 1mg if the uncertainty in its velocity is 5.5×10^{-20} ms^{-1}.

8. On the basis of Heisenberg uncertainty principle, show that electron cannot exist within the atomic nucleus (Radius = 10^{-15} m). h = 6.626×10^{-34} Jsec.

9. Why can't we overcome the uncertainty predicted by Heisenberg's principle by building more precise devices to reduce the error in measurement below the h/4π limit?

10. What physical meaning is attributed to the square of the absolute value of wave function $|\Psi|^2$?

11. Which of the four quantum numbers (n, l, m_l, m_s) determine (a) the energy of an electron in a hydrogen atom and in a many-electron atom. (b) The size of an orbital. (c) the shape of an orbital, (d) the orientation of an orbital in space?

12. Draw the shapes (boundary surfaces) of the following orbitals : (a) $2p_y$, (b) $3d_{z^2}$, (c) $3d_{x^2-y^2}$. (Show coordinate axes in your sketches).

13. Discuss the similarities and differences between a 1s and a 2s orbital.

14. For each of the following pair of hydrogen orbitals, indicate which is higher in energy:
    (A) 1s, 2s;   (B) 2p, 3p;   (C) $3d_{xy}$, $3d_{yz}$;   (D) 3s, 3d;   (E) 4f, 5s

15. Which orbital in each of the following pairs is lower in energy in a many-electron atom?
    (a) 2s, 2p;   (b) 3p, 3d;   (c) 3s, 4s;   (d) 4d , 5f

16. Explain the meaning of the symbol $4d^6$.

17. The ground-state electron configurations listed here are incorrect. Explain what mistakes have been made in each and write the correct electron configurations.
    Al $1s^2 2s^2 2p^4 3s^2 3p^3$ B $1s^2 2s^2 2p^5$ F $1s^2 2s^2 2p^4$

18. Draw orbital diagrams for atoms with the following electronic configuration:
    (a) $1s^2 2s^2 2p^3$ (b) $1s^2 2s^2 2p^6 3s^2 3p^5$ (c) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^7$

19. What do you understand by
    (i)   radial probability density, R^2 and
    (ii)   radial probability function. $4\pi r^2 R^2$? How do they vary with r for $1s$, $2s$ and $2p$ atomic orbitals of hydrogen atoms.

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