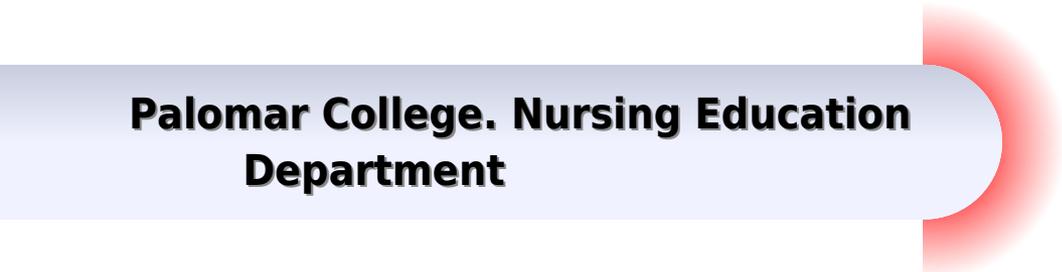




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web 1. molarity  $M = \frac{\text{mass}}{\text{molar mass} \times \text{volume}}$  4500 g / (108.25 g/mol  $\times$  0.0129 m<sup>3</sup>) = 34.8 mol/L  
 0.0581158 mol / 0.00581158 mol/L = 10 kg  
 0.0129 m<sup>3</sup> mole fraction NaOH = 4.825 g / (141.9579 g/mol  $\times$  0.00581158 mol/L) = 0.4500 g / 18.015 g/mol = 24.97918401 mol

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web apr 7 2014 molarity moles of solute / liters of solution 8.422 a first convert 250 ml to liters 250 / 1000 = 0.25 then calculate molarity 5 moles / 0.25 liters = 20 M  
 c a solution with molarity 2 requires 2 M of NaOH per liter so 4  $\times$  2.8 M = 11.2 M  
 a solution of molarity 1.5 M requires 1.5 mol of NaOH to every litre of solvent

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web answer the correct option is c explanation the molarity of a mixture can be calculated using the following formula  $M_{\text{mix}} = \frac{M_1 V_1 + M_2 V_2}{V_1 + V_2}$  1.5  $\times$  480 + 2  $\times$  520 / (480 + 520) = 1.824 M  
 q5 with increase in temperature which of these changes a molality b fraction of solute present in water

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