



Galaxy Note 10guide

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2 p 1 b 1 s 2 2 s 1 2 p 1 1 s 2 2 p 2 c 1 s 2 2 p 2 1 s 2 1 p 2 d 1 s 2 1 p 2

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web elements in the second row of the periodic table place their electrons in the 2n shell as well as the 1n shell for instance lithium li has three electrons two fill the 1 s orbital and the third is placed in the 2 s orbital giving an electron configuration of 1 s 2 2 s 1

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web predictions from reliable sources have been used for these elements grayed out electron numbers indicate subshells filled to their maximum bracketed noble gas symbols on the left represent inner configurations that are the same in each period written out these are he 2 helium 1s 2 ne 10 neon 1s 2 2s 2 2p 6

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potassium element has 19 electrons its electron configuration is $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$ or simply put $4s^1$ what is the total number of electrons that can be filled in s p and d subshells

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