

Holt Modern Chemistry Review

CHAPTER 6: CHEMICAL BONDING

The following pages contain the bulk (but not all) of the information for the chapter 6 test. Focus on this content, but make sure to review class notes, activities, handouts, questions, etc. If you study this document and NOTHING else, you should at least be able to PASS the test. ***** Test items will be recall, examples, and/or application of this content. *****

OUTCOMES

- Collaborate with peer(s) to understand chemistry content (C C)
- Communicate chemistry content to teacher and peer(s) (E C)
- 6.2: Use the Lewis dot notation, illustrate the covalent bond as the sharing of electrons (F & PK)
- 6.3: List important physical properties of ionic compounds; sum of oxidation numbers = zero (F & PK)

6.1: INTRODUCTION TO CHEMICAL BONDING

- **Vocabulary**
 - **chemical bond** -- the attractive force that holds atoms or ions together
 - **ionic bonding** -- a force that attracts electrons from one atom to another, which transforms a neutral atom into an ion
 - **covalent bonding** -- a bond formed when atoms share one or more pairs of electrons
 - **nonpolar-covalent bond** -- a covalent bond in which the bonding electrons are equally attracted to both bonded atoms
 - **polar** -- describes a molecule in which the positive and negative charges are separated
 - **polar-covalent bond** -- a covalent bond in which a pair of electrons shared by two atoms is held more closely by one atom
- **Chapter Highlights**
 - Most atoms are chemically bonded to other atoms. The three major types of chemical bonding are ionic, covalent, and metallic.
 - In general, atoms of metals bond ionically with atoms of nonmetals, atoms of metals bond metallically with each other, and atoms of nonmetals bond covalently with each other.

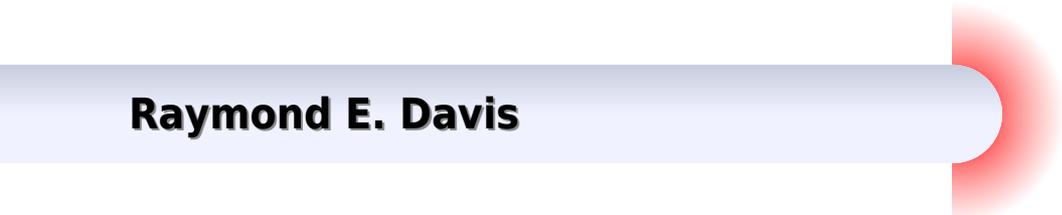
6.2: COVALENT BONDING AND MOLECULAR COMPOUNDS

- **Vocabulary**
 - **molecule** -- a group of atoms that are held together by chemical forces; a molecule is the smallest unit of matter that can exist by itself and retain all of a substance's chemical properties
 - **molecular compound** -- a chemical compound whose simplest units are molecules
 - **chemical formula** -- a combination of chemical symbols and numbers to represent a substance
 - **molecular formula** -- a chemical formula that shows the number and kinds of atoms in a molecule, but not the arrangement of the atoms
 - **electron-dot notation** -- an electron configuration notation in which only the valence electrons of an atom of a particular element are shown, indicated by dots placed around the element's symbol
 - **Lewis structure** -- a structural formula in which electrons are represented by dots; dot pairs or dashes between two atomic symbols represent pairs in covalent bonds.
See examples at right for LEWIS STRUCTURES of NF₃ and OCl₂
 - **structural formula** -- a formula that indicates the location of the atoms, groups, or ions relative to one another in a molecule and that indicates the number and location of chemical bonds
 - **single bond** -- a covalent bond in which two atoms share one pair of electrons
 - **multiple bond** -- a bond in which the atoms share more than one pair of electrons, such as a double bond or a triple bond
 - **resonance** -- the bonding in molecules or ions that cannot be correctly represented by a single Lewis structure
- **Chapter Highlights**
 - Atoms in molecules are joined by covalent bonds. In a covalent bond, two atoms share one or more pairs of electrons.



Holt Chemistry Chapter 6

Raymond E. Davis



Holt Chemistry Chapter 6:

Holt Chemistry Ralph Thomas Myers, 2004 **Holt Chemistry File**, 1998 This reference is a must for students who need extra help reteaching or extra practice The guide moves students through the same concepts as the text but at a slower pace More descriptive detail along with visual algorithms provides a more structured approach Each chapter closes with a large bank of practice problems Book jacket *Forthcoming Books* Rose Arny, 2003 Books in Print Supplement, 2002

The Chemistry of the Nitro and Nitroso Groups Henry Feuer, 1969 Nanodiamond Oliver A Williams, 2014-03-18 The exceptional mechanical optical surface and biocompatibility properties of nanodiamond have gained it much interest Exhibiting the outstanding bulk properties of diamond at the nanoscale in the form of a film or small particle makes it an inexpensive alternative for many applications Nanodiamond is the first comprehensive book on the subject The book reviews the state of the art of nanodiamond films and particles covering the fundamentals of growth purification and spectroscopy and some of its diverse applications such as MEMS drug delivery and biomarkers and biosensing Specific chapters include the theory of nanodiamond diamond nucleation low temperature growth diamond nanowires electrochemistry of nanodiamond nanodiamond flexible implants and cell labelling with nanodiamond particles Edited by a leading expert in nanodiamonds this is the perfect resource for those new to and active in nanodiamond research and those interested in its applications *Adhesives Technology Handbook* Sina Ebnesaajad, Arthur H. Landrock, 2008-09-16 Following the successful first the second edition is a complete guide to all that is required to successfully bond materials It is both a reference and a source for learning the basics for those involved in the entire product value chains Basic principles of adhesion such as surface characterization types of adhesive bonds and adhesion failure topics are covered in addition to a description of common adhesive materials and application techniques Provides the end user practitioners of adhesion technology with a complete guide to bonding materials successfully Covers most substrates including plastics metals elastomers and ceramics explaining basic principles and describing common materials and application techniques Arranges information so that each chapter can be studied selectively or in conjunction with others **Handbook of Adhesives and Surface Preparation** Sina Ebnesaajad, 2010-12-15 Handbook of Adhesives and Surface Preparation provides a thoroughly practical survey of all aspects of adhesives technology from selection and surface preparation to industrial applications and health and environmental factors The resulting handbook is a hard working reference for a wide range of engineers and technicians working in the adhesives industry and a variety of industry sectors that make considerable use of adhesives Particular attention is given to adhesives applications in the automotive aerospace medical dental and electronics sectors A handbook that truly focuses on the applied aspects of adhesives selection and applications this is a book that won't gather dust on the shelf Provides practical techniques for rendering materials surfaces adherable Sector based studies explore the specific issues for automotive and aerospace medical dental and electronics **KY HS Test Prac Wkbks W/Corr Sci 2001** Holt

Rinehart & Winston,2002-03 *Children's Books in Print, 2007 ,2006* **Comprehensive Chemical Kinetics** Clement Henry Bamford,Charles Frank Howlett Tipper,1973 Contemporary Organic Chemistry Andrew L. Ternay,1976

Chemistry for the Life Sciences J. G. Dawber,Alan Thomas Moore,1973 Chemistry For the Life Sciences covers the basic chemistry of HNC HND courses in applied biology and medical laboratory subjects but the treatment is also appropriate to post HNC courses in biochemistry such as the M I Biol and to certain degree level courses The text is designed to provide a unified approach to the subject with roughly equivalent amounts of physical and organic chemistry and a small amount of inorganic chemistry in chapter one It proceeds from a relatively pure chemistry in the early chapters to aspects of biochemistry and molecular biology in later chapters Publisher *The Science Teacher* ,1979 Some issues are accompanied by a CD ROM on a selected topic **Comprehensive Chemical Kinetics: Organic reactions. v. 8. Proton transfer** C. H. Bamford,Charles Frank Howlett Tipper,1973

Trace Environmental Quantitative Analysis Paul R. Loconto,2020-12-27 A thorough and timely update this new edition presents principles techniques and applications in this sub discipline of analytical chemistry for quantifying traces of potentially toxic organic and inorganic chemical substances found in air soil fish and water as well as serum plasma urine and other body fluids The author addresses regulatory aspects calibration verification and the statistical treatment of analytical data including instrument detection limits quality assurance quality control sampling and sample preparation and techniques that are used to quantify trace concentrations of organic and inorganic chemical substances Key Features Fundamental principles are introduced for the more significant experimental approaches to sample preparation Principles of instrumental analysis determinative techniques for trace organics and trace inorganics analysis An introduction to the statistical treatment of trace analytical data How to calculate instrument detection limits based on weighted least squares confidence band calibration statistics Includes an updated series of student tested experiments *An Introduction to Atmospheric Physics* Robert G. Fleagle,Joost A. Businger,1981-01-09

This book is addressed to those who wish to understand the relationship between atmospheric phenomena and the nature of matter as expressed in the principles of physics The interesting atmospheric phenomena are more than applications of gravitation of thermodynamics of hydrodynamics or of electrodynamics and mastery of the results of controlled experiment and of the related theory alone does not imply an understanding of atmospheric phenomena This distinction arises because the extent and the complexity of the atmosphere permit effects and interactions that are entirely negligible in the laboratory or are deliberately excluded from it the objective of laboratory physics is by isolating the relevant variables to reveal the fundamental properties of matter whereas the objective of atmospheric physics or of any observational science is to understand those phenomena that are characteristic of the whole system For these reasons the exposition of atmospheric physics requires substantial extensions of classical physics It also requires that understanding be based on a coherent way of seeing the ensemble of atmospheric phenomena Only then is understanding likely to stimulate still more general insights

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Thompson, Richard Salter Storrs, Joshua Leavitt, Henry Ward Beecher, Theodore Tilton, Henry Chandler Bowen, William Hayes
Ward, Hamilton Holt, Fabian Franklin, Harold de Wolf Fuller, Christian Archibald Herter, 1913 *Children's Books in Print* R R
Bowker Publishing, Bowker, 1999-12

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