

Lab 22A

Investigating Chemical Equilibrium

Purpose: To investigate the macroscopic properties of two chemical systems at equilibrium, observe shifts in equilibrium concentrations as a function of concentration and temperature and relate the to systems, LeChatelier's principle and equilibrium constant applying Le Chatelier's principle.

Proble:

1. Write a balanced equation for the reaction in part 1.
$$\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$$
2. Write a balanced equation for the reaction in part 2.
$$\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$$
3. Write a balanced equation for the reaction in part 3.
$$\text{FeSCN}^{2+}(\text{aq}) + \text{H}^{+}(\text{aq}) \rightleftharpoons \text{FeSCNH}^{3+}(\text{aq})$$
4. In part 4, what is the conjugate acid of the base FeSCN^{2+} ? Write the K_{a} expression for the acid. What is the conjugate base of the acid? Write the K_{b} expression for the base. How are the K_{a} and K_{b} related? How do you predict the direction of the reaction? How do you predict the direction of the reaction? How do you predict the direction of the reaction?
5. What is the conjugate acid of the base FeSCN^{2+} ? Write the K_{a} expression for the acid. What is the conjugate base of the acid? Write the K_{b} expression for the base. How are the K_{a} and K_{b} related? How do you predict the direction of the reaction? How do you predict the direction of the reaction? How do you predict the direction of the reaction?
6. What is the conjugate acid of the base FeSCN^{2+} ? Write the K_{a} expression for the acid. What is the conjugate base of the acid? Write the K_{b} expression for the base. How are the K_{a} and K_{b} related? How do you predict the direction of the reaction? How do you predict the direction of the reaction? How do you predict the direction of the reaction?
7. Write the equation for the reaction between Fe^{3+} and OH^{-} ions.
$$\text{Fe}^{3+}(\text{aq}) + \text{OH}^{-}(\text{aq}) \rightleftharpoons \text{Fe}(\text{OH})^{2+}(\text{aq})$$
8. Write the equation for the reaction between Fe^{3+} and OH^{-} ions.
$$\text{Fe}^{3+}(\text{aq}) + \text{OH}^{-}(\text{aq}) \rightleftharpoons \text{Fe}(\text{OH})^{2+}(\text{aq})$$
9. Write the equation for the reaction between Fe^{3+} and OH^{-} ions.
$$\text{Fe}^{3+}(\text{aq}) + \text{OH}^{-}(\text{aq}) \rightleftharpoons \text{Fe}(\text{OH})^{2+}(\text{aq})$$

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