

# HSC Chemistry Notes:

## Production of Materials:

### 1. Fossil Fuel products

#### 1.0. Construct Word and Balanced Formulae Equations of Chemical Equations that are encountered.



#### 1.1 Identify the Industrial Source for Ethylene from the cracking of some of the fractions from the refining of petroleum.

Oil refineries are required to balance their petroleum outputs to the demands of the marketplace, i.e. more gasoline produced for cars rather than grease. This doesn't occur in the natural process of fractional distillation and hence catalytic cracking is applied, converting lower demanded fractions into gasoline and ethylene (ethene).

**Catalytic Cracking:** High boiling point (molecular weight) fractions from crude oil are broken into lower boiling point (molecular weight) substances to increase the output of high demand products.



**Zeolites**, which are crystalline aluminosilicates, are used as a surface catalyst for catalytic cracking. Its ordered structure and large surface area allows it to effectively begin the 'cracking' process in its inner surface to outer, in 500 degrees plus and above atmospheric pressure, creating ethene and propene as by-products.

**Steam Cracking or Thermal cracking** is a non-catalyst cracking process where alkanes are passed extremely hot metal tubes, decomposing them into ethene, propene or butane.

#### 1.2 Identify that ethylene because of the high reactivity of its double bond is readily transformed into many useful products.

Most simple alkane and ethene are very similar in characteristics, they are both gases at room temperature and low boiling points, held together only by dispersion forces. However due to alkene's double bond in contrast to alkane's single bond, alkene's become more reactive by **opening out the double bond to form two single bonds, known as addition reactions.**

Alkenes such as ethene react with chlorine, bromine and iodine, decolorizing them.

**Note: Alkyl Groups** are alkane groups with one hydrogen missing.



Ethene can readily transform into useful products such as ethanol with the reaction of water and phosphoric acid as a catalyst.

Ethylene can also readily react with oxygen via a silver catalyst forming ethylene oxide (sterilizer) which then may be converted into ethylene glycol used as polymers and automotive antifreeze.

Can also be polymerized to polyethylene (plastic), and acetic acid (food preservative).

# Hsc Chemistry Notes

**Angela Brew, Judyth Sachs**



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*HSC Chemistry Notes*, 1984      **HSC Chemistry Complete Course Notes**, 2016-06-01      **HSC Year 12 Chemistry Notes (2019)**, 2018      Lecture/tutorial Notes and Problems for an Extension Course in H.S.C. Chemistry, 1981

**Lecture Notes in Chemistry I. (Harvard College.)** Arthur Prescott Lothrop, Archibald Livingston Hodges, 1882

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**Redirecting Science** Finn Aaserud, 2003-01-30 This volume is an important study for understanding the complex interconnections between basic science and its sources of economic support in the period between the two world wars The focus of the study is on the Institute for Theoretical Physics later renamed the Niels Bohr Institute at Copenhagen University and the role of its director the eminent Danish physicist Niels Bohr in the funding and administration of the Institute Under Bohr's direction the Copenhagen Institute was a central workplace in the development and the formulation of quantum mechanics in the 1920s and later became an important center for nuclear research in the 1930s Dr Aaserud brings together the scholarship on the internal origins and development of nuclear physics in the 1930s with descriptions of the concurrent changes in private support for international basic science particularly as represented by Rockefeller Foundation philanthropy In the process the book places the emergence of nuclear physics in a larger historical context This book will appeal to historians of science physicists and advanced students in these areas

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